



## Synthesis, Evaluation of asymmetric dicationic Ionic Liquids and studying their efficiency for Egyptian Base Oil improvement

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### Abstract

In this work, three asymmetric dicationic ionic liquids, namely, 1-(2-(1-decyl-1*H*-imidazolium-3-yl)ethyl)-4-methylpyridinium bromide (IL<sub>101</sub>), 1-(5-(1-decyl-1*H*-imidazolium-3-yl)pentyl)-4-methylpyridinium bromide (IL<sub>102</sub>), and 1-(10-(1-decyl-1*H*-imidazolium-3-yl)decyl)-4-methyl-pyridinium bromide (IL<sub>103</sub>) were synthesized. Their structures were confirmed and characterized via, elemental analysis, Fourier transform infra-red spectroscopy and proton nuclear magnetic resonance. The surface properties of these ionic liquids, surface tension, critical micelle concentration, effectiveness, maximum surface excess, and minimum surface area, were investigated with different concentrations at 25°C. Standard free energies of micellization and adsorption were obtained. They were studied and tested as base oil improvers, especially oxidation stability, (through studying the change in total acid number [TAN] and viscosity), and foaming characteristics. The efficiency of these compounds as base oil improvers is ranked as follows, IL<sub>103</sub> > IL<sub>102</sub> > IL<sub>101</sub>.

**Keywords:** Ionic liquid; Surface parameters; Base oil; Oxidation stability; Total acid number; Foaming, Spectroscopy.

### 1. Introduction

In lubricant applications there is always a drive to reduce emissions, increase durability, and increase service intervals, as well as to reduce friction and thus energy use. To meet these challenges, lubricants must be always improved. These improvements can also lead to technological advances by enabling the use of new materials for a particular application [1].

Generally, there are three different base oils used in the formulation of lubricating oil, these are; mineral, bio-based (natural) and synthetic base oils [2, 3].

Mineral base oils used in engineering machines were derived from one of the distillation fractions of crude oil containing (C20-C50) hydrocarbons. However, the base oil will be predominantly iso-paraffinic and naphthenic hydrocarbons if the fraction boiling is under 723 K together with a few percent of mono aromatic hydrocarbons [3, 4].

Lubricants based on biological sources are able to adhere to metal surfaces, enabling improved boundary lubrication. In a contrary, their drawbacks are associated with their high sensitivity to oxidation,

hydrolytic instability and high pour points [3, 5-7].

Many synthetic lubricants have been produced and some of them have been applied such as poly-alpha olefins (PAOs). Many examples of synthetic oils are; synthetic hydrocarbon polymers, poly-phenyl ethers, halogenated hydrocarbons, carboxylate esters, aliphatic esters, polyol esters, phosphate esters and poly-alkylene glycols [8-11].

The prolonged use of lubricants alone on the contact surfaces results in ramification to corrosion and wear, which directly affects the engine efficiency. Oxidation is considered as the most important chemical phenomenon in lubricating oils [12, 13]. Total Acid Number (TAN) is one of tools used for measuring of acidic matter remaining in the oil because of the degradation process that occurred. Also, foam control is a very important factor in froth flotation of ores and other minerals, and the mechanisms of foam destruction in froth flotation are similar to those encountered in antifoam applications [14, 15].

Ionic liquids (ILs) are a class of organic molten salts consisting of only cations and anions [16]. ILs have extensive applications because of their unique

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physicochemical characteristics such as negligible vapor pressure, high thermal stability, larger surface tensions, higher shear viscosities, high thermal conductivity, wide electrochemical window and low melting point, etc. [17]. Compared to traditional monocationic ionic liquids, dicationic ionic liquids (DILs), especially asymmetrical dicationic ionic liquids (ADILs), have shown to possess superior physicochemical properties in the field of volatility and thermal stability [18]. Therefore, they can be used as for many applications including dye-sensitized cells, organic synthesis, anti-corrosive additives, coatings or antioxidants, lubricants and lubricant additives [19].

Findings from several studies conducted in the past years say that the addition of engine oil additives act as defoamers [14,15,20], viscosity improvers, pour-point depressants, anti-wear additives, extreme pressure additives, etc. Thus, engine oil additives could serve as potential solution to the problem of reduced efficiencies [14].

The use of additives is very important. Their selection is based on the ability of them to reduce friction and wear, controlling viscosity, improving viscosity index, resist corrosion and oxidation [15,20-22], increase oil lifetime through prevention, and monitoring of oil degradation. Not only the oil quality is the main consideration but also the economic considerations are important. e.g., the cost of oil can be very expensive for the huge machinery that holding several thousand liters of lubricating oil [21]. Different examples of common additives include dispersants, antioxidants, friction modifiers, anti-wear agents, detergents and viscosity index (VI) improvers. The most available constituents of modern engine oils are discussed in many literatures [23- 25].

In our previous work we synthesized and used different heterocyclic compounds, [26-32], azo-phenols, [33] and azo-liquid crystals [34-36] for base oil improvement such as antioxidants, viscosity improvers and pour point depressants.

In this present work we synthesized and studied the effect of some asymmetric dicationic ionic liquids as antioxidants, viscosity improvers and anti-foams for tested Egyptian hydro finished base oil.

## 2. Experimental

### 2.1. Materials and Methodology

i- physicochemical characteristics of delivered Base oil, (Co-operation Petroleum Company) It is carried out according to ASTM standard test methods, Table 1.

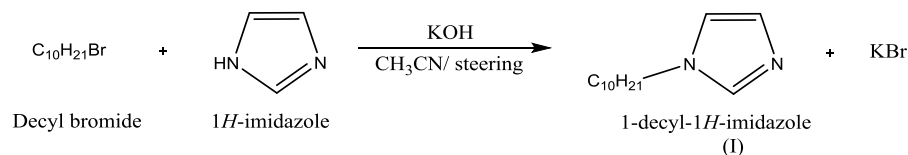
ii- Chemicals and solvents, purchased from international chemical companies, were used without further purification. Acetonitrile (anhydrous), ethyl acetate (99.5%), absolute ethyl alcohol, 1-bromodecane (98%), 1H-imidazole (99%), petroleum ether (40–60) (98%), and potassium hydroxide (97%), Merck. 4-Methyl pyridine (99%), 1,2-dibromoethane (99%), 1,5-dibromopentane (99%), 1,10-dibromodecane (99%), Alfa Aesar.

iii- Synthesis of dicationic ionic liquids IL<sub>101-103</sub>

Steps of synthesis were illustrated as follows:

Compound (I) was prepared previously [37]. In a conical flask, 1H-imidazole (0.1 mol) and potassium hydroxide (0.1 mol) were dissolved in acetonitrile (50 ml) with stirring at room temperature. 1-Bromodecane (0.11 mol) was added dropwise for 1h. Stirring is continued till white precipitate of potassium bromide is formed. The precipitate was eliminated by filtration and the filtrate is evaporated under vacuum. The product is washed several times with ethyl acetate, to remove the excess reactants, and evaporated under vacuum equation 1.

1,10-dibromodecane) (0.015 mol) at room temperature for 24 hrs. The mixtures were washed several times with ethyl acetate to remove unreacted reactants.



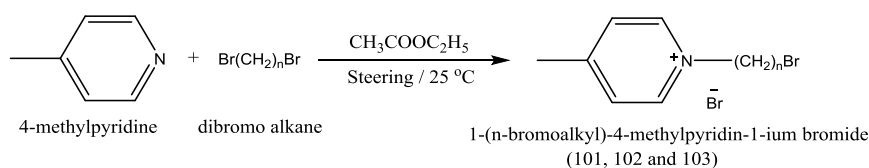
**Equation 1: Preparation of 1-decyl-1H-imidazole (I)**

**Table 1 Physicochemical properties of the Base oil**

Test	Result	Standard Test Method
Density @ 15.56 °C, g /cm <sup>3</sup>	0.8807	ASTM D – 1298
Pour Point, °C	0	ASTM D – 97
Viscosity @ 40 °C	165.33	ASTM D – 445
@ 100 °C	15.56	ASTM D – 445
Viscosity Index (VI)	95	ASTM D – 2270
Total Acid Number (TAN)	0.061	ASTM D – 664
Foaming Characteristics at 50°C,		ASTM D – 892
Foaming Volume (ml)	170	
Collapse Time (s)	480	
Sulfur Content, (wt.%)	0.41	ASTM D – 4294
Color	3.5	ASTM D – 1500
Copper Corrosion	1a	ASTM D – 130
Flash Point, °C	220	ASTM D – 92
Ash Content, (wt.%)	0.003	ASTM D – 482
Wax Content, wt.%	1.6	UOP - 46
Saturate Content, wt.%	54.36	-----
Aromatic Content, wt.%	45.54	-----
Molecular weight	321.23	-----

Compounds [101, 102 and 103] were prepared by stirring 4-methyl pyridine (0.01 mol) and dibromo alkanes, (1,2-dibromoethane, 1,6-dibromohexane and

White precipitate was afforded by filtration and recrystallized from petroleum ether 40–60, equation 2, [38].

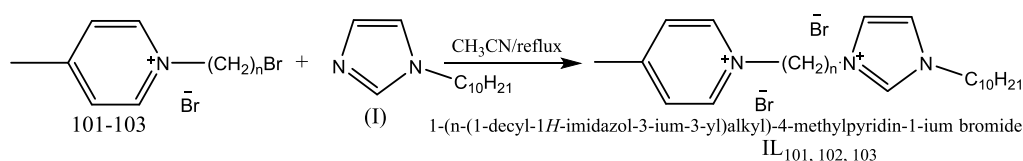


### Equation 2: Preparation of 1-(n-bromoalkyl)-4-methylpyridin-1-ium bromide

101; n = 2, 1-(2-bromoethyl)-4-methylpyridin-1-ium bromide  
 102; n = 5, 1-(5-bromopentyl)-4-methylpyridin-1-ium bromide  
 103; n = 10, 1-(10-bromodecyl)-4-methylpyridin-1-ium bromide

Synthesis of Ionic Liquids IL<sub>101-103</sub>: These compounds were synthesized by refluxing compound (I) (0.01 mol.) with compounds 101-103, (0.01 mol.) in 50 ml acetonitrile for 12 h at 80 °C. Purification of

Metrohm Titrand 905 potentiometer, stirrer 801 and electrode model 6.0229.010. Instrument parameters used according to Metrohm Application Bulletin AB-404/2e.



### Equation 3: Preparation of IL<sub>101-103</sub>

IL<sub>101</sub>: 1-(2-(1-decyl-1H-imidazol-3-ium-3-yl) ethyl)-4-methylpyridin-1-ium bromide.  
 IL<sub>102</sub>: 1-(5-(1-decyl-1H-imidazol-3-ium-3-yl) pentyl)-4-methylpyridin-1-ium bromide.  
 IL<sub>103</sub>: 1-(10-(1-decyl-1H-imidazol-3-ium-3-yl) decyl)-4-methylpyridin-1-ium bromide

The products was carried out via evaporation of acetonitrile under vacuum, equation 3, [39].

The synthesized compounds were characterized using the conventional tools of analysis, Elemental

analysis (Micro analytical center, Cairo, Egypt using Elementary Viro El Microanalysis), FT-IR spectrophotometer (Model 960 Mooog, ATI Mattson Infinity Series, USA, Cairo, Egypt) and 1H-NMR spectra (Bruker High Performance Digital 1H-NMR

Spectrometer Advance III 400 MHz using DMSO as solvent.

### 2.2. Surface tension measurements:

Surface tension measurements were obtained using a Du Noüy tensiometer with a platinum ring. Freshly prepared aqueous solutions of asymmetric IL<sub>101-103</sub> bola surfactants were measured over a concentration range of 0.01–0.000001 M/L at 25°C. The surface tension of double distilled water was measured to calibrate instruments, which was generally 72:00 ± 0:50 mN/m. Every sample solution was prepared with the double distilled water which was stabilized for 5 min in the instrument before measurements and was repeated three times to reduce the error [40,41].

### 2.3. Total Acid Number:

This test underwent for the base oil without and with the additives via ASTM D664 [42], using

### 2.4. Viscosity:

According to ASTM D445 [43], viscosities of the samples without and with the additives were

measured using CT-1000 CANON oil bath, Cannon-Fenske viscometers size 150, and Kohler thermometer instrument.

### 2.5. Foaming test:

foaming characteristics were measured for the blank sample and the three samples after adding the additives using ASTM D892 [44] (Petrotest DP water bath and ZEAL DM3A gas meter)

## 3. Results and discussion

### 3.1. Characterization of IL<sub>101-103</sub>

#### 3.1.1. Elemental analysis:

Data tabulated in Table 2 revealed that the calculated values are in good matching with the observed ones.

#### 3.1.2. FT-IR:

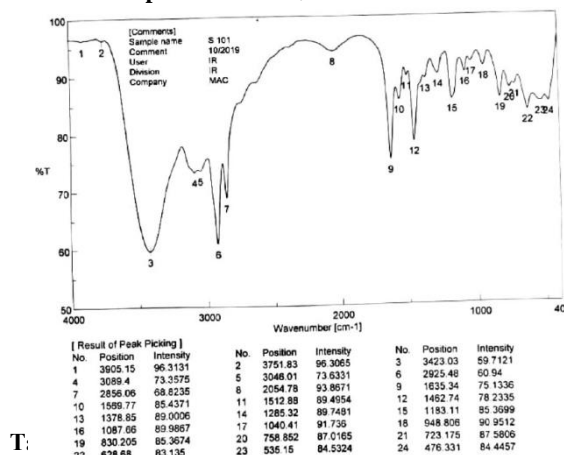
Data recorded in table (3) and charts 1-3 illustrated the following:

**Table 2: Elemental analysis of the ionic liquids (IL<sub>101-103</sub>).**

IL#	C%		H%		N%		Br%		Mol. Wt.
	Calc.	Obs.	Calc.	Obs.	Calc.	Obs.	Calc.	Obs.	
IL <sub>101</sub>	51.55	51.34	7.21	7.19	8.59	8.64	32.66	32.82	489
IL <sub>102</sub>	54.24	54.08	7.78	7.82	7.91	8.05	30.07	29.35	531
IL <sub>103</sub>	57.90	57.68	8.55	8.63	6.99	7.04	26.57	26.64	602

- In the FT-IR spectra of compounds (IL<sub>101-103</sub>), the aromatic stretching vibrations bands of benzene C-H bonds are observed in the range of 3128-3039 cm<sup>-1</sup>. Stretching vibrations bands of aromatic C=C are observed in the regions of 1512, 1511 and 1517 cm<sup>-1</sup> respectively.
- N-H broad stretching bands, appeared at 3423, 3425 and 3412 cm<sup>-1</sup>, for (IL<sub>101-103</sub>) are attributed to the presence of the carbene proton in the form of NH<sup>+</sup> in imidazole ring [35].
- Aromatic C-H stretching bands appeared at 3089-3046, 3090(broad), and 3128-3039 cm<sup>-1</sup> respectively.
- Aliphatic C-H (stretching) bands appeared at 2925-2856, 2924-2854 and 2926-2856 cm<sup>-1</sup>.
- All the specific imidazole zone bands were obtained precisely [39].

**Chart 1: IR spectrum of IL<sub>101</sub>**



IL	(v cm <sup>-1</sup> )											
	NH+	CH Ar	CH Aliphatic	C=C Ar	C-C Ar	C-C C-N	C-H in plane bending	C-H out of plane bending	Ring deformation in plane bending	Ar bending zone	Ring deformation out of plane	
IL 101	3423 broad	3089-3046	2925, 2856	1512	1462	1569 1462-1378	1285-1183	758	1040-948	830, 758	628	
IL 102	3425 broad	3090 broad	2924-2854	1511	1462	1565 1462-1373	1292-1231	731	1081-913	830, 731	626	
IL 103	3412 broad	3128-3039	2926-2856	1517	1463	1564 1463-1372	1308-1240	754	1044-931	831-754	631	

Chart 2: IR spectrum of IL102

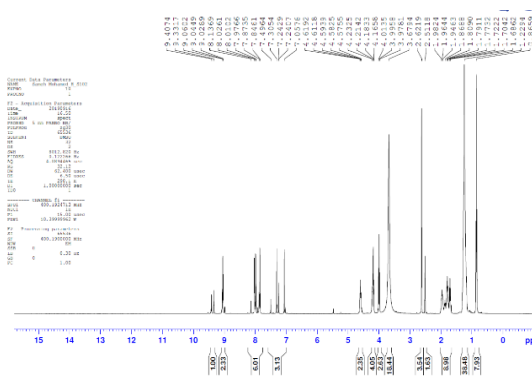
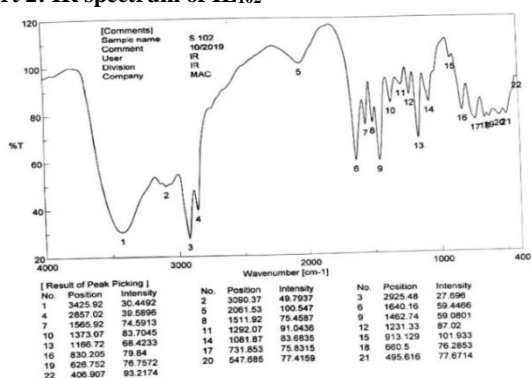


Chart 3: IR spectrum of IL103

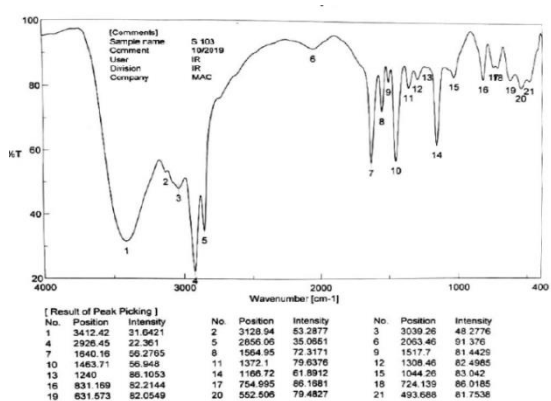


Chart 6: NMR spectrum of IL103

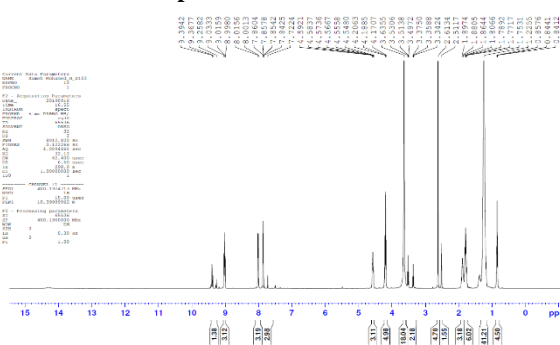
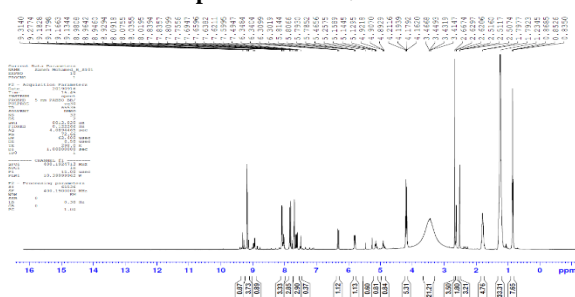
3.1.3. <sup>1</sup>H-NMR:

Table 4 and charts 4-6 shows the obtained results of <sup>1</sup>H-NMR for prepared compounds IL<sub>101-103</sub>.

Chart 4: NMR spectrum of IL101



## 3.2. Surface activity

The critical micellar concentrations (cmc) determined from the surface tension measurements are given in Table 5.

The surface excess concentration ( $\Gamma_{max}$ ) and minimum area per molecule at the air-water interface ( $A_{min}$ ) determined from surface tension data using Equations (i and ii) [45-47] are given in Table 5.

$$\Gamma_{max} = -1/nRT (d\gamma/d\ln C) \quad (i)$$

$$A_{min} = 10^{16} / N_A \Gamma_{max} \quad (ii)$$

Where R is gas constant, T is temperature,  $\gamma$  is the surface tension, and  $N_A$  is Avogadro's number. The constant n takes the values 2 for an ionic surfactant where the surfactant ion and the counter ion are univalent and n = 3 for a dimeric surfactant made up

of a divalent surfactant ion and two univalent counter ions, in the absence of a swamping electrolyte [46]. Hence a value of  $n = 3$  was used in Eq. (i) for the present ionic dimeric surfactant system.

The results of surface tension measurements are shown in Fig. 2. Table 5 lists the values of CMC, surface tension at CMC ( $\gamma_{CMC}$ ),  $\Gamma_{max}$ ,  $A_{min}$ , and  $P_{C20}$  of IL<sub>101-103</sub>.

As listed in Fig. 2, these asymmetric ionic liquids have the same head groups and hydrophobic chains, but differs in spacer chain lengths. The spacer should, Therefore, be the principal factor determining differences in surface properties.

As shown in Table 5, these compounds have low CMC values. They formed compact micelles due to strong hydrophobic interaction between multiple hydrocarbon chains. We noticed that, with the increase in spacer chain length,  $\gamma_{CMC}$  does not show a marked difference with change in the spacer length ( $m = 2, 5$  or  $10$ ), and the CMC values decrease for the di cationic ionic liquids.

The area occupied per molecule ( $A_{min}$ ) decreased with the increase in spacer chain length, which will improve the flexibility of the molecular and reduce the backbone rigid and pack more tightly at the

air/water contact, which causes the decrease in  $A_{min}$ . The spacer may penetrate the hydrophobic core. The saturation surface excess concentration ( $\Gamma_{max}$ ) increases as  $A_{min}$  decreases. Due to the decrease in repulsion between the oriented ionic heads at the interface when the spacer bend more to the hydrophobic phase, adsorbing as a folded, wicket-like conformation [48].

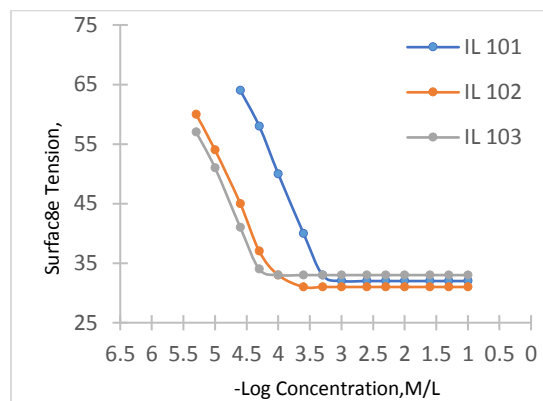


Figure 1: Surface tension isotherms of IL<sub>101-103</sub> at 25 °C

Table 4: Chemical Shift ( $\delta$  ppm) for compound IL<sub>103-105</sub>.

Compound	Proton type ( $\delta$ ppm)											
	a	b	c	d	e	f	g	h	I	j	k, k'	L
IL <sub>101</sub>	9.31 (s)	9.16 (d)	8.08 (d)	7.83 (d)	7.59 (d)	5.25 (t)	4.19 (t)	2.67 (s)	1.79 (m)	1.23 (s)	0.86	9.31 (s)
IL <sub>102</sub>	9.33 (s)	9.04 (d)	8.02 (d)	7.97 (d)	7.24 (d)	4.59 (t)	4.21 (t)	2.62 (s)	1.96 (m)	1.79 (m)	1.70 (m) 1.22 (s)	0.84 (t)
IL <sub>103</sub>	9.36 (s)	9.01 (d)	8.00 (d)	7.86 (d)	7.75 (s)	4.58 (t)	4.18 (t)	2.61 (s)	1.86 (m)	1.78 (m)	1.76 (m) 1.23 (s)	0.84 (t)

Table 5: Surface parameters of the synthesized surfactants [IL<sub>101-103</sub>]

Compound	CMC mol./L	$\gamma_{CMC}$ mN/m	$\pi_{CMC}$ mN/m	$P_{C20}$	$\Gamma_{max} \times 10^{10}$ mol./cm <sup>2</sup>	$A_{min}$ nm <sup>2</sup>	$\Delta G^{\circ}_{mic}$ KJ/mol	$\Delta G^{\circ}_{ads}$ KJ/mol
IL <sub>101</sub>	$6 \times 10^{-4}$	32	40	$7 \times 10^{-5}$	1.359	1.222	-18.381	-47.824
IL <sub>102</sub>	$1 \times 10^{-4}$	31	41	$1.2 \times 10^{-6}$	1.372	1.210	-22.823	-52.703
IL <sub>103</sub>	$5 \times 10^{-5}$	33	39	$8 \times 10^{-6}$	1.452	1.144	-24.540	-51.412

The difference between the surface tension of IL<sub>101-103</sub> at their CMC and that of pure water is termed "effectiveness" ( $\pi_{CMC}$ ):  $\pi_{CMC} = \gamma_0 - \gamma$

where  $\gamma_0$  is the surface tension of the pure water and  $\gamma$  is the surface tension of the solution at CMC [49]. The effectiveness of the IL<sub>101-103</sub> ranged between

39 and 41 dyne/cm at 25°C (Table 5). The most efficient surfactant was the one that lowered the surface tension at the CMC. It is illustrated that these compounds are efficient in achieving the maximum reduction of the surface tension at CMC.

Efficiency ( $P_{C20}$ ): The  $P_{C20}$  values were determined in Table (5).  $P_{C20}$  is the bulk surfactant

concentration needed to reduce the surface tension of the solvent by an arbitrary 20 mNm<sup>-1</sup>.

This value commonly characterized the efficiency of a surfactant to lower surface tension. Values of C20 showed that IL<sub>101-103</sub> indicated that they have great efficiency in reducing surface tension of water and consequently, have high surface activity. Standard free energy of micellization ( $\Delta G^{\circ}_{mic}$ ), and standard free energy of adsorption ( $\Delta G^{\circ}_{ads}$ ):

$$\Delta G^{\circ}_{mic} \text{ is calculated as } \Delta G^{\circ}_{mic} = RT \ln CMC.$$

$\Delta G^{\circ}_{mic}$  for the synthesized IL<sub>101-103</sub> surfactant were always negative values, indicating that micellization was a spontaneous process. The general trend of  $-\Delta G^{\circ}_{mic}$  increase with increasing the distance between the two heads groups (spacer), i.e., it became less negative. This indicates that increasing in the spacer length and hydrophobicity favor the micellization process.

**Table 6: Variation of total acid number (mg KOH/g) without and with IL<sub>101-103</sub>**

Sample	ILs Conc. ppm	Time intervals (hrs.)		without	and with
		24	48		
Base oil					
Blank (without additive)	0.061	0.2621	0.4636	0.7056	0.9130
IL <sub>101</sub>	200	0.1935	0.3145	0.3752	0.4032
	400	0.0826	0.1895	0.2670	0.3545
	600	0.0408	0.0605	0.0934	0.1747
IL <sub>102</sub>	200	0.1915	0.2580	0.3225	0.3896
	400	0.0625	0.1613	0.2419	0.3230
	600	0.0403	0.0504	0.0835	0.1512
IL <sub>103</sub>	200	0.0615	0.1126	0.1895	0.2923
	400	0.0346	0.0406	0.0797	0.1613
	600	0.0302	0.0346	0.0635	0.0775

Conversely,  $\Delta G^{\circ}_{ads}$  can be calculated by the relation

$$\Delta G^{\circ}_{ads} = \Delta G^{\circ}_{mic} - (0.6 \times \pi_{CMC} \times A_{min})$$

Table 5 shows some increase in  $\Delta G^{\circ}_{ads}$ , supporting the idea of micellization over adsorption on the solution surface to overcome the repulsion forces occurring at the water/hydrophobe interface.  $\Delta G^{\circ}_{ads}$  and  $\Delta G^{\circ}_{mic}$  were always negative in value due to the spontaneity, and indicating that the processes are thermodynamically favored [50, 51].

### 3.3. Effect of IL<sub>101-103</sub> as base oil improvers:

We studied the effect of these prepared compounds as base oil improvers via:

#### 3.3.1. Improvement in Total Acid Number values (TAN):

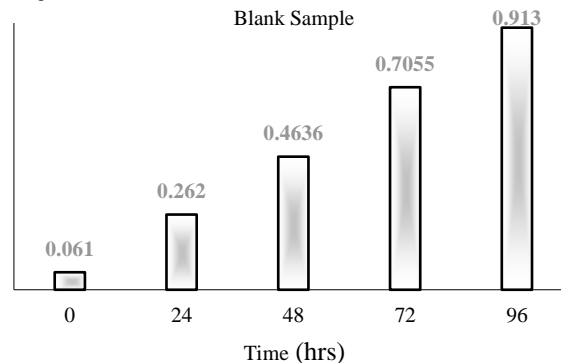
Total acid number is used as a guide in the quality control of lubricating oil. It is also used as a measure of lubricant degradation in service [52]. Figure 2 illustrated the behavior of blank sample against the oxidation process (from zero time up to 96 h, at 120 °C). It is obvious that, the rate of oxidation is increased with time. This result is attributed to that, oil is possibly being contaminated with acidic constituents, increasing the acid content in the oil, Scheme 1.

The maximum total acid number value appeared after 96 h (table 5). We noticed that, after adding compounds (IL<sub>101-103</sub>) with different concentrations, the values of total acid number decreased.

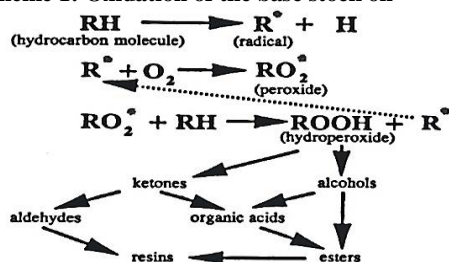
Table 6 and figures (3-5) revealed that after thermal oxidation of the base oil up to 96 hours, in presence of additives, (IL<sub>101-103</sub>), the TAN values decreased with increasing additives concentrations which ranged from 0.1935 up to 0.4032 mg KOH/g (200 ppm), 0.0826 to

0.3545 mg KOH/g (400 ppm), and from 0.0408 to 0.1747 mg KOH/g (600 ppm) respectively, Figures (3-5). This means that, stability of the base oil is increased by adding the prepared compounds, where the efficiency is ranked as follows: IL<sub>103</sub> > IL<sub>102</sub> > IL<sub>101</sub>.

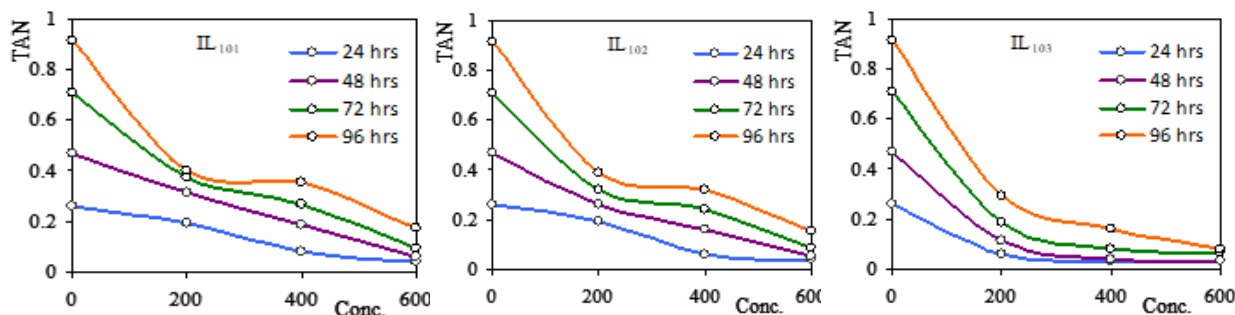
**Figure 2: Variation of TAN values of blank sample with mg KOH/g**



## Scheme 1: Oxidation of the base stock oil



Figures 3, 4 and 5: TAN (mg KOH/g) variation of the base oil with and without IL<sub>101</sub>, IL<sub>102</sub>, and IL<sub>103</sub> (respectively) of concentrations 200, 400, and 600 ppm



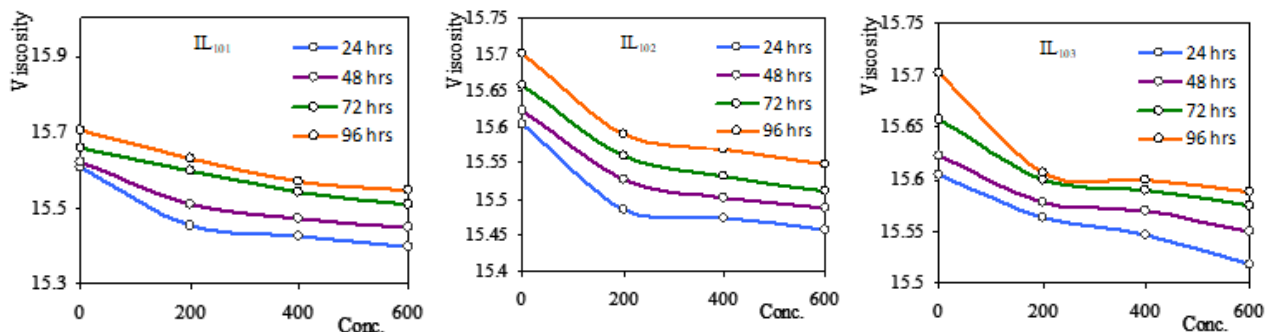
and sludge, Table 7. The blank sample gave high reading (15.70 mm<sup>2</sup>/s) in viscosity after oxidation for 96 h. The viscosity values showed good results after adding the prepared compounds with different concentrations, table 7 and figures (6-8). These results illustrated that; the oxidation stability is improved in presence of the prepared additives.

### 3.3.3. Effect of IL101-103 on foaming characteristics

Foaming is a common problem with oil-lubricated components. Foam is a collection of small bubbles of air that accumulate on or near the surface of the fluid. The presence of air bubbles in the fluid can lead to excessive oxidation, cavitation, the reduction of

characteristics where the obtained foaming volume is considered high (170 ml) with long time (480 s) for foaming collapse, table 8. As shown in Table 8 and figures 9a-11b, the foaming characteristics of the base, after adding the prepared compounds with different concentrations were reduced compared to the blank sample, (from 480 up to 65 ml). Also, the foaming collapse time was reduced from 480.03 to 120.15 s. The efficiency of the additives as antifoams is ranked as follows: IL<sub>103</sub> > IL<sub>102</sub> > IL<sub>101</sub>.

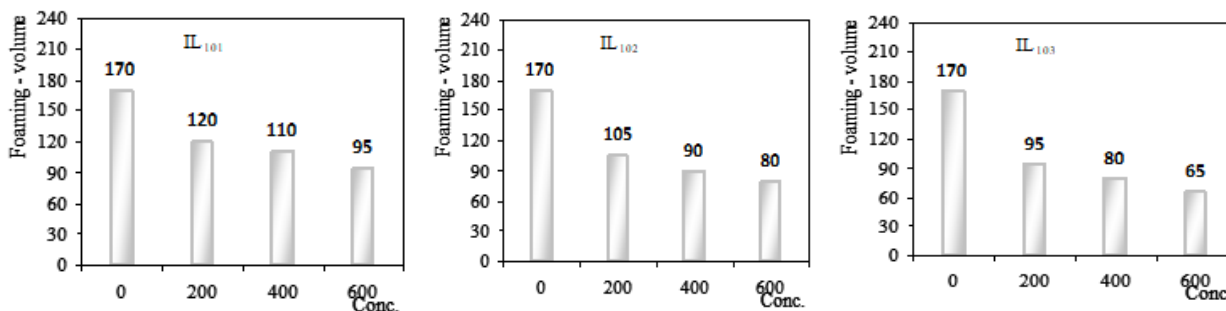
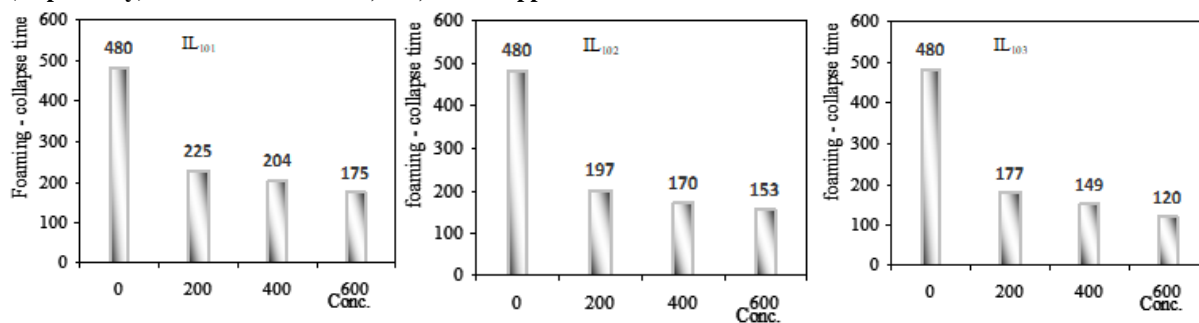
Figures 6, 7 and 8: Viscosity (mm<sup>2</sup>/s) variation of the base oil with and without IL<sub>101</sub>, IL<sub>102</sub>, and IL<sub>103</sub> (respectively) of concentration: 200, 400, and 600 ppm





**Table 7: Variation of viscosity at 100<sup>o</sup> C (mm<sup>2</sup>/s) with and without IL101-103**

Sample	ILs Conc. ppm	Time intervals (hrs.)			
		24	48	72	96
Base oil Blank (without additive)	0	15.60	15.62	15.66	15.70
IL <sub>101</sub>	200	15.46	15.51	15.60	15.63
	400	15.42	15.47	15.54	15.57
	600	15.40	15.45	15.51	15.55
IL <sub>102</sub>	200	15.49	15.53	15.56	15.59
	400	15.48	15.50	15.53	15.57
	600	15.46	15.49	15.51	15.55
IL <sub>103</sub>	200	15.56	15.58	15.60	15.61
	400	15.55	15.57	15.59	15.60
	600	15.52	15.55	15.58	15.59

**Figures 9, 10, and 11(a): Foaming volume (ml) variation of the base oil with and without IL<sub>101</sub>, IL<sub>102</sub>, and IL<sub>103</sub> (respectively) of concentrations 200, 400, and 600 ppm****Figures 9, 10, and 11 (b): Foaming collapse time (s) variation of the base oil with and without IL<sub>101</sub>, IL<sub>102</sub>, and IL<sub>103</sub> (respectively) of concentrations 200, 400, and 600 ppm**

**Table 8: Variation of foaming volume (ml) at 50 °C with and without additives**

Sample	ILs Conc. ppm	Foaming Volume (ml)	Foaming Collapse Time (s)
Base oil Blank (without additive)	0	170	480.03
IL <sub>101</sub>	200	120	225.23
	400	110	204.09
	600	95	174.89
IL <sub>102</sub>	200	105	197.39
	400	90	169.55
	600	80	153.45
IL <sub>103</sub>	200	95	176.92
	400	80	149.29
	600	65	120.15

#### 4. Conclusions

\* The physicochemical characteristics of the delivered base oil is described.

\* The structures of the prepared three compounds (IL<sub>101-103</sub>) were well elucidated.

\* Their surface parameters indicated that, they have good surface activity.

\* They were tested as antioxidants and antifoams for local base oil, and the efficiency of them is ranked as follows: IL<sub>103</sub> > IL<sub>102</sub> > IL<sub>101</sub>.

#### Conflicts of interest

There are no conflicts to declare.

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