

## Corrosion Study of Glaze-Ceramics Doped Cement-Kiln-Dust Soaked in Concentrated HCl

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**N**EW glaze-ceramics formulated from industrial waste have been prepared. The method consists of introducing cement-kiln-dust (CKD) in the industrial glaze-ceramics with various ratios (10-50%) together with the addition of 10gm boric acid. The chemical durability of the prepared composite cement-dust-glaze-ceramics (CDGC) is investigated using high concentrations of (2,5 & 8N) HCl acid solutions. CDGC are vitrified by melting for four hours at 1400°C ± 250°C depending on the amount of CKD introduced. Specifically, the experimental results indicate that the introduction of CKD in industrial glaze decreases durability. At high HCl concentrations, the durabilities of the prepared glaze-ceramics are observed to decrease with increasing the percentage of (CKD) until certain limit and almost constancy. The corrosion results are explained according to the ion exchange mechanism and formation or dissociation of the modified depleted layer.

Some hollow and jogged areas are observed by SEM which might be interpreted as evidence of local corrosion phenomena. In addition, the pH values of the leachant solutions are measured. The comparison of hardness for the optimum glaze-ceramic and glaze-ceramic derivatives prepared in this study indicates a degradation % in their chemical and physical properties depending on the amount% of CKD introduced. It is observed that the degradation not only depends on the amount of CKD introduced, but also on the strength of the leaching HCl acid solution. Also, the results show that, the prepared glazes have a surface hardness values ranged from 5.5 to 2.5 in the Mohs scale.

**Keywords:** CKD waste, New glaze-ceramics, Corrosion and Hardness.

The re-use of materials gained from solid wastes reduces degradation of living environment, which is the initial objective of solid-waste management. In this context, the minimization of huge wastes produced by the cement industry is of great concern in most of the countries. For a sustainable future, it would be imperative to recycle as much as of these wastes as possible, on one hand to alleviate the problem of environmental inconvenience caused by landfills and on

the other hand to create a supply of inert materials that would meet the environmental laws and industrial policies<sup>(1)</sup>.

Ceramics have come up as a trend to reduce the risks to human and the environment as well as offering the possibility of combining the special properties of conventional sintered glaze-ceramics with the distinctive characteristics of cement-kiln-dust waste. Developing ceramic-glazes demonstrates the advantage of combining various remarkable properties in one material. Ceramic-glazes comprise special low melting type of silicate glasses. For each type of ceramic bare product, it is necessary to develop the corresponding glazes because of the need to match the chemical and physical properties. About 500 000 tons of (CKD) wastes are generated during the industrial production of cement. At present, the waste is being deposited in open area where it pollutes the environment. Its utilization will result in significant cost but also the problem of pollution of the environment will be solved.

Glasses are generally chemically stable both in acidic and basic environments. This characteristic feature of glass made vitrification techniques useful for treatment of different wastes<sup>(2)</sup>. However, this method leaves vitreous product requiring disposal. Although these products are chemically inert, it is important to evaluate chemical durability. Exposure of glass to water may at first leach mobile modifier ions followed by the dissolution of glass matrix. These corrosion reactions start from the surface and proceed inwards by diffusion of various ionic species as well as by dissociation of chemical bonds which form glasses. The overall reaction involves two subsequent primary steps; leaching of mobile ions and uniform (or complete) dissolution of the matrix itself<sup>(3)</sup>.

For silicate glasses, alkali leaching kinetics are often described by two limiting stages. In the first stage, alkali leaching proceeds as a function of  $t^{1/2}$ . In the most commonly used leaching model, this stage is attributed to a simple ion exchange process in which  $H^+$  or  $H_3O^+$  from the leachate diffuses into the glass, and alkali cations diffuse out. The rate of interdiffusion is often equated with the self-diffusion coefficient of the slowest moving cation in bulk glass. The common explanation for the second stage of leaching, where alkali leaching is linear with time, is that the leached layer eventually becomes as thick as the dissolution rate of the silicate network, resulting in uniform glass dissolution, and a leached layer which reaches a constant steady state thickness as a function of time<sup>(4)</sup>.

This paper presents results regarding the corrosion behavior of the CDGC soaked in high concentrated HCl solution. In this work, chemical durability and hardness properties of glazes doped CKD derivatives have been studied with the focus on the principal role of CKD to modifying the chemical durability of the commercial glaze and topography of the surface. Finally, this study aims to know how the addition of CKD to the industrial glaze enables the CDGC to resist high concentrated HCl solutions.

## Experimental

### Glaze preparation

Tables 1 and 2 show the composition of CKD and prepared industrial glazed-ceramics, respectively. CKD used in the experiments was sampled from Beni-Suef Company, while glaze was from Shiny Company. The chemical compositions for both CKD and industrial glaze were analyzed with a combination of inductively coupled plasma (ICP) and X-ray fluorescence spectroscopy (XRF). The new glaze batches were formulated as shown in Table 3. The cement dust was added with increasingly amounts in the different batches from 10% to 50%. The batches were melted individually in platinum crucibles 2%Rd at ~1400-1150°C for 5 hr depending on the CKD% and were stirred frequently for complete homogeneity until reasonably free from air bubbles.

**TABLE 1. Analysis of portland cement and CKD (Wt.%).**

Constituents	CaO	SiO <sub>2</sub>	Al <sub>2</sub> O <sub>3</sub>	Fe <sub>2</sub> O <sub>3</sub>	MgO	K <sub>2</sub> O	Na <sub>2</sub> O	SO <sub>3</sub>	Cl	I.L.
Portl. Cem.	63.77	20.36	5.33	3.97	2.1	0.14	0.48	2.72	0.03	19.10
CKD	45.14	11.45	3.22	1.88	0.53	3.73	2.72	5.14	2.59	23.60

**TABLE 2. Frit composition for the industrial glaze (weight%).**

Borax	B <sub>2</sub> O <sub>3</sub>	Albite	ZnO	Dolomite	CaCO <sub>3</sub>	Zircon	Sand	Na <sub>2</sub> (CO <sub>3</sub> )
11.418	27.71	12.026	0.78	3.634	8.992	13.929	18.788	2.722

**TABLE 3. Cement-dust-glaze (CDG) compositions in weight%.**

Glaze notation	Frit	B <sub>2</sub> O <sub>3</sub>	CKD
G1	90	10	0
G2	80	10	10
G3	70	10	20
G4	60	10	30
G5	50	10	50

Slabs with dimensions 1 x 1 x 0.4 cm<sup>3</sup>, were shaped in a stainless steel mold and selected on the basis of uniformity of dimensions and freedom from bubbles. The specimens were annealed in an electric furnace at 500°C and after soaking for 1 hr were left to cool slowly over-night at a rate of 20°C/hr.

### Chemical resistance test

A series of glaze coupons (1x1x0.2 cm) was placed horizontally in 50 ml polyethylene BD Falcon™ 50 ml Conical Tube with Flip Top Cap which allows the etchant solution (50ml) to completely cover the whole glass surfaces of the specimen and then the tube is tightly closed. The inlet acid solutions (2,5 & 8 N HCl) used in this study were prepared from deionized water plus reagent grade of HCl acid. The test duration ranged from 1 up to 25 days at room temperature (~25°C) and, at 100°C± 2°C from 2 up to 30 hr. After complete specified immersion times in each acid solution, samples were cleaned in de-ionized water and left to dry completely in air in a desiccator and then accurately weighed. As the etching is affected by surface area, so the relationship between the geometric area of the glass samples and the volume of the solutions was maintained as S/V= 0.042 cm<sup>2</sup>/ml.

*pH measurements*

Outlet fluid pH was measured immediately after corrosion tests using an Orion Research (601A) at 25°C. Accuracy of measurements was  $\pm 0.01$ pH unit. Initial (before experiments) solutions were used as standards to minimize matrix defect.

*SEM*

Photomicrographs of CDGC samples before and after the corrosion experiments were performed using a JSM- 5400 Scanning electron microscope (SEM). A thin film of gold was previously evaporated on the surface before measurement.

*Hardness test (Mohs scratch test)*

Mohs scale of mineral hardness characterizes the scratch resistance of the material in this case the glazed surface. The Mohs scale is based on ten minerals that are all readily available except the last one, diamond (Table 4). The hardness of a material is measured against the scale by finding the softest/hardest mineral that can scratch the material. On the Mohs scale, value between 1-5 are considered soft surface while those values above 5 up to 10 are considered hard surface.

**TABLE 4. Hardness of minerals based on Mohs scale .**

Mohs Hardness	Minerals	Absolute hardness
1	Talc ( $Mg_3Si_4O_{10}(OH)_2$ )	1
2	Gypsum ( $CaSO_4 \cdot 2H_2O$ )	3
3	Calcite ( $CaCO_3$ )	9
4	Fluorite ( $CaF_2$ )	21
5	Apatite ( $Ca_5(PO_4)_3(OH, Cl, F)$ )	48
6	Orthodase Feldspar ( $KAlSi_3O_8$ )	72
7	Quartz ( $SiO_2$ )	100
8	Topaz ( $Al_2SiO_4(OH, F)_2$ )	200
9	Corundum ( $Al_2O_3$ )	400
10	Diamond ( C )	1500

The glazed-tile test pieces (1x2x0.4cm) were cleaned and dried prior to testing. The prepared test pieces were subjected to scratch using the mineral found in the Mohs scratch test kit. The scratch while using as minimal pressure as possible was done starting with the mineral of lowest hardness value, which is 1 on the Mohs scale and subsequently changed to the next mineral that could inflict or cause scratch to the glazed-tile surface.

**Results***Leaching*

The degree of corrosion in a glaze specimens is calculated from its weight loss ( $gm/cm^2$ ) over time (days). Figures 1-5 show the weight loss for the five studied glazes (G1- G5), respectively as a function of time (1-25 days) when immersed in 2, 5 & 8N HCl solutions at room temperature. Overall weight losses of the glaze are observed to be less in 2N HCl solution compared to the weight loss in 5 or 8N HCl solutions. Also, it can be noticed that, the weight losses

increase drastically either with increasing immersion times or HCl concentrations especially at early immersion times (1-10 days). However, with increasing immersion times, there is a swinging in the weight losses depending on the CKD% as shown in Fig. 3-5. The effect of temperature (100°C) is presented in Fig.6 and it is evident that the corrosion increases relatively to the increase of CKD% (up to 40%) and then either decreases or remains nearly unchanged (50% CKD) when the glaze samples are immersed in 2 or 5N HCl, respectively. The effect of immersion times (6-30 hrs) for glazes 3 and 5 when soaked in 2 or 5N HCl are illustrated in Fig. 7 and 8, respectively. From these figures it can be noticed that the corrosion rate is slow at the early immersion times (6-18 hr) and then becomes faster afterwards.

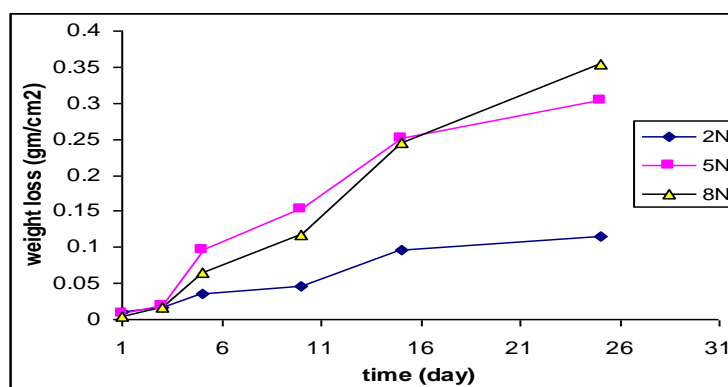


Fig. 1. Effect of immersion time (days) and HCl concentrations on the durability of glaze (G1) at room temperature.

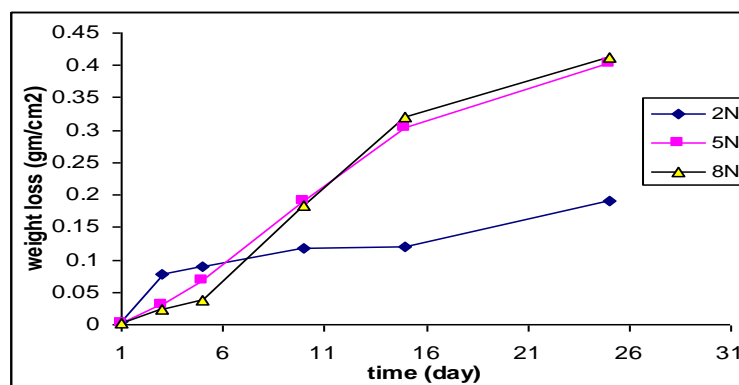


Fig. 2. Effect of immersion time (days) and HCl concentrations on the durability of glaze (G2) at room temperature.

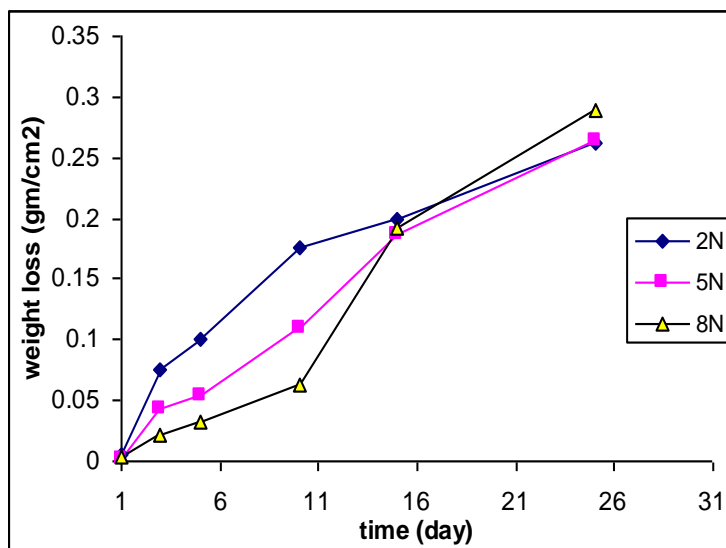


Fig. 3. Effect of immersion time (days) and HCl concentrations on the durability of glaze (G3) at room temperature.

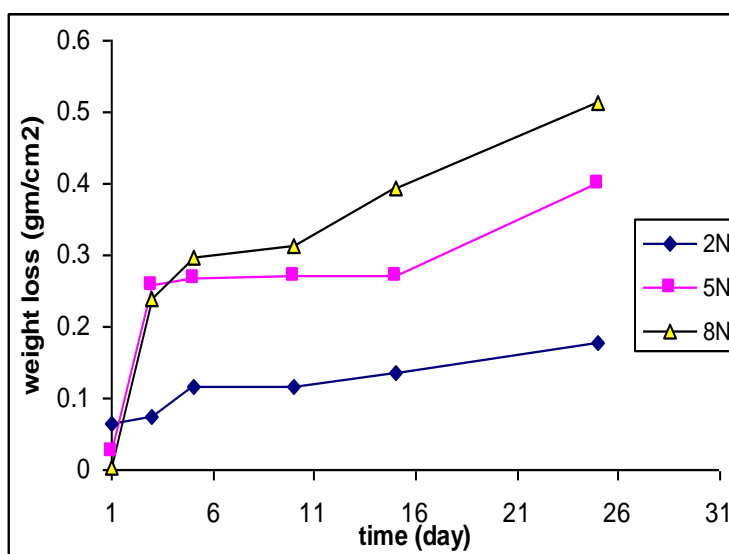


Fig. 4. Effect of immersion time (days) and HCl concentrations on the durability of glaze (G4) at room temperature.

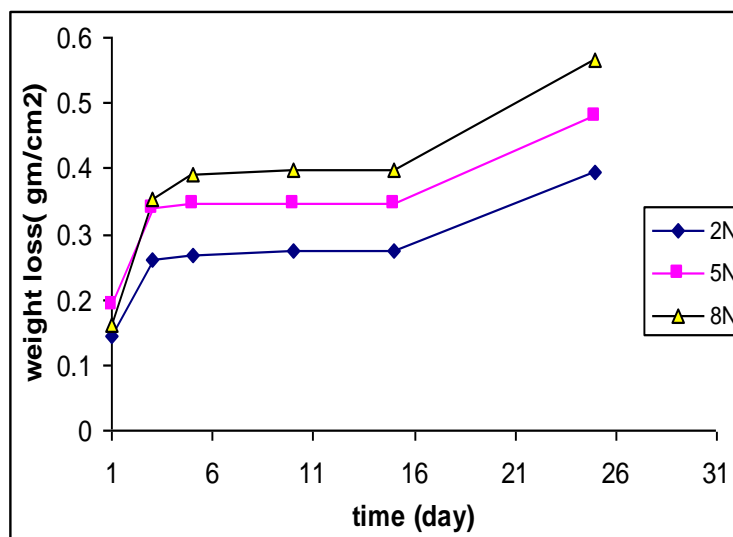


Fig. 5. Effect of immersion time (days) and HCl concentrations on the durability of glaze (G5) at room temperature.

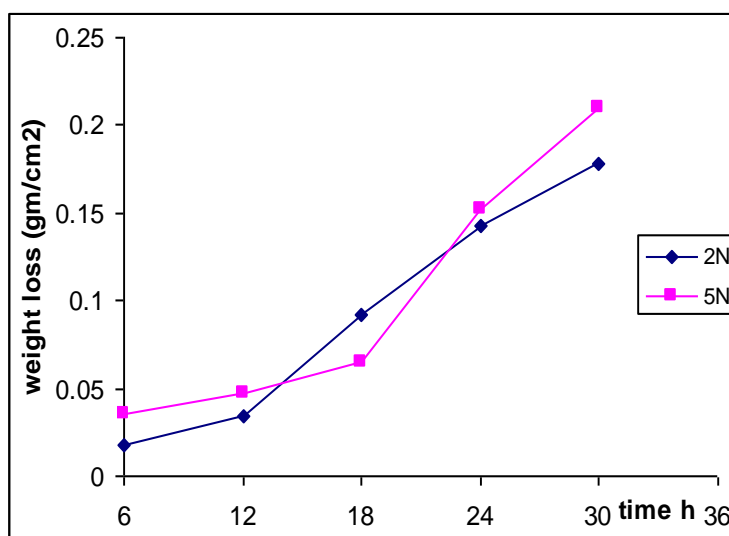


Fig. 6. Effect of immersion time (hours) and 2,5N HCl concentrations on the durability of glaze (G3) at 100°C.

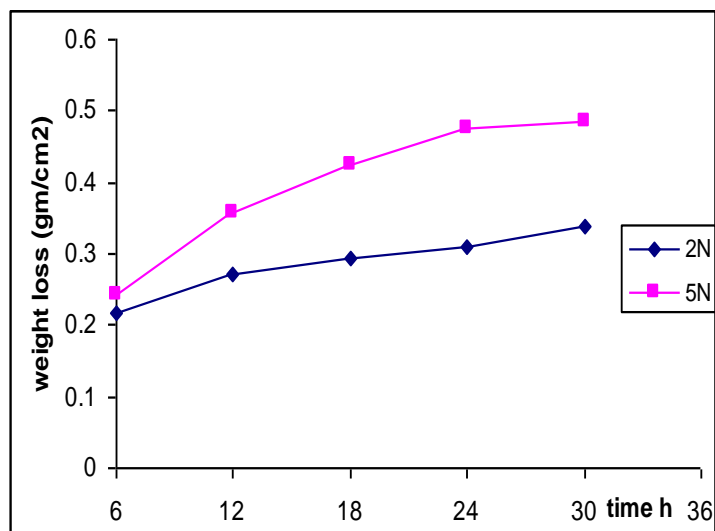


Fig. 7. Effect of immersion time (hours) and 2,5N HCl concentrations on the durability of glaze (G5) at 100°C.

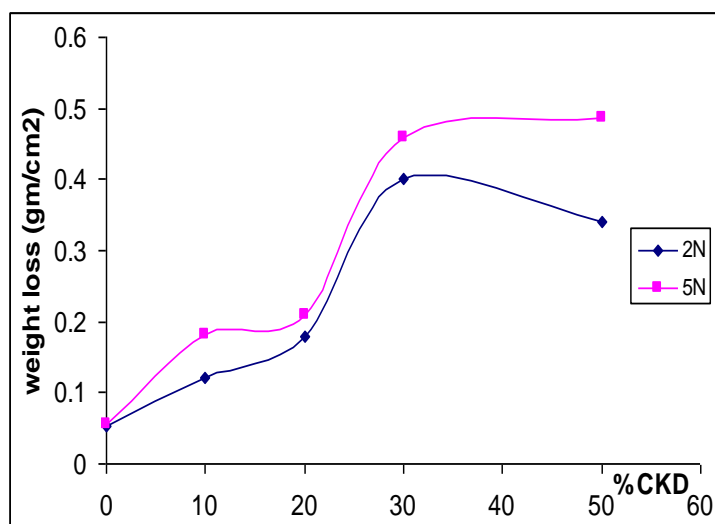


Fig. 8. Effect of glaze compositions and 2,5N HCl concentrations on the durability of glaze (G3) at 100°C.



*pH measurements*

Figure 9 reveals that the pH values for the leachate relatively increase as the CKD% or the strength of HCl solutions increases. Moreover, the same trend is observed when the temperature of HCl solutions is raised to 100°C or the strength of HCl solution is decreased (Fig. 10).

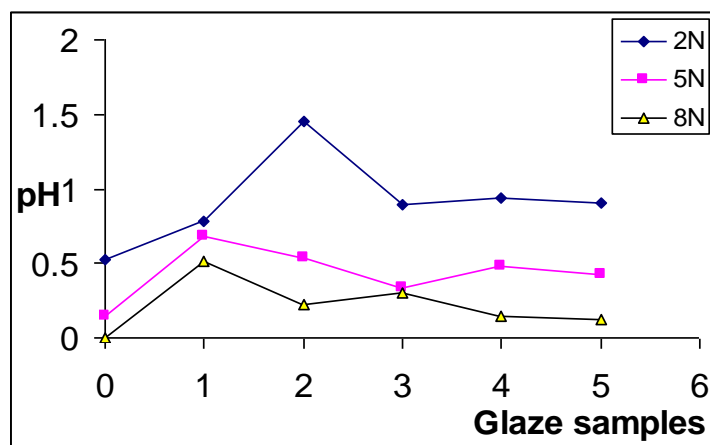


Fig. 9. Effect of glaze compositions and 2,5,8N HCl concentrations on the pH value at room temperature.

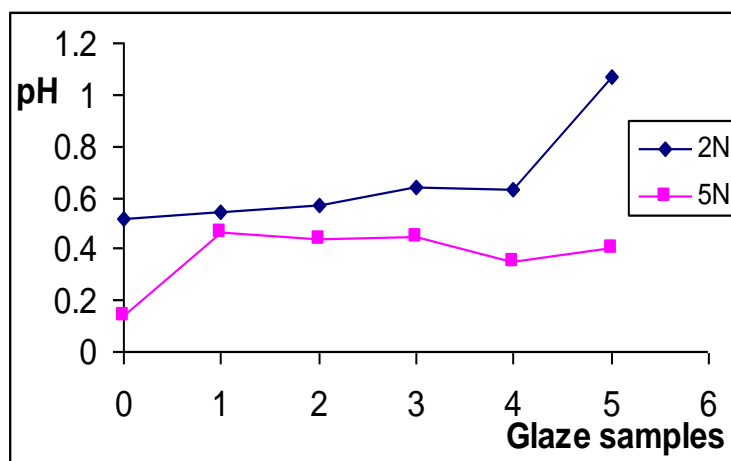


Fig. 10. Effect of glaze compositions and 2,5,8N HCl concentrations on the pH value at 100°C.

### *Hardness*

Figures 12-15 show the results of Mohs scratch test on the 5 prepared CDGC samples after been immersed in different HCl acid solutions for different immersion times. All the five specimens were found to be scratched by the mineral fluorite, which on the Mohs scale has a hardness of 6 (Table 4). The values of absolute hardness for the corroded glaze samples (G1,G2, G3) are illustrated in Fig. 12 -14. It can be noticed that, the hardness generally decreases gradually with increasing HCl concentrations or the immersion times. The decrease is sharply in the early stages of corrosion and then there is some sort of stability. In addition, Fig. 15 shows the effect of different CKD% on the hardness values after been soaked in 5N HCl solution for 25 days. It is obvious that, the hardness values fall down rapidly for G5 at the early immersion times (up to 10 days) followed by a constancy in the hardness values except that, the decrease in hardness values is related oppositely to the CKD%. Generally, in our case the prepared glazes have a surface hardness  $< 5.5$  and  $> 2.5$  which are considered to be normal and soft surfaces.

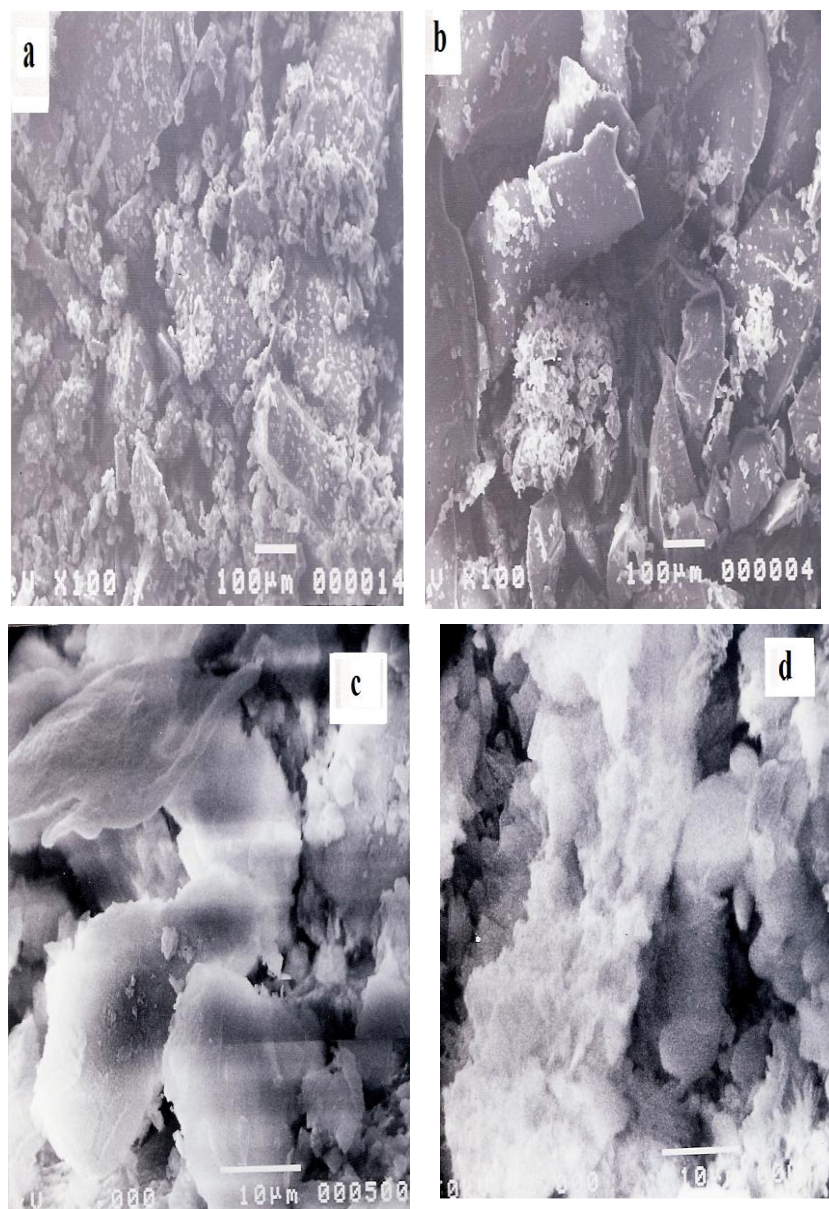
### **Discussion**

Due to the rapid increase in the quantities of solid wastes, affecting both public health and environment, efficient and sustainable solid waste-management must be considered as the most pressing issue on the agenda of developing countries and governments worldwide.

### *Corrosion*

Chemical durability of glazes is usually discussed in terms of durability of the surface glassy phase in accordance with the same terms as with that occurring with glasses. Glasses react in acidic environment by ion exchange or preferentially substitution of alkali ions by hydrogen ions. In alkaline environments, the network structure of glass is continuously attacked and destroyed by hydroxyl attack<sup>(5-9)</sup>. Glaze-ceramic materials are considered as polycrystalline silicate solids containing a glassy phase and possess a valuable combination of the favorable properties of both glasses and ceramics<sup>(10)</sup>.

This influence of different oxides on the durability of glasses and glossy glazes is well understood. However, chemical resistance of glazes consisting of one or several crystalline phases embedded in a glassy matrix has not been fully and widely studied. Wollastonite and anorthite crystals devitrified from a fritted glaze have been reported to be attacked by acidic solutions<sup>(11)</sup>. Wollastonite crystals in fast-fired raw glazes are attacked by acidic and also slightly alkaline water solutions<sup>(12,13)</sup>. The formation of crystalline phases in traditionally fired glazes takes place according to the equilibrium reactions and is controlled by the total oxide composition of the glaze. However, in a modern fast-firing process of floor tiles, the short firing cycle of 60-90min restricts the extent of raw materials<sup>(14)</sup>. The surface is often unmaturing and its chemical and mechanical properties are incompletely understood<sup>(15)</sup>.



**Fig. 11.** SEM of glaze G3 before and after soaking in 2,5,8N HCl concentrations for 25 days at room temperature.

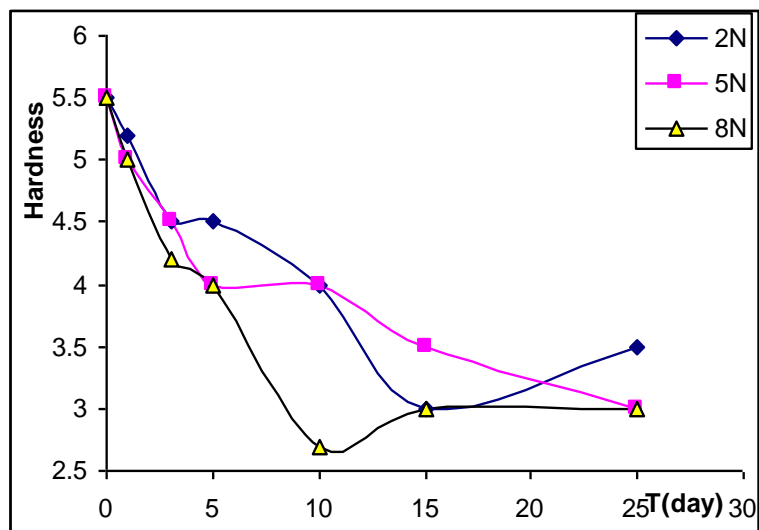


Fig. 12. Effect of immersion time (days) and HCl concentrations on the absolute hardness of glaze (G1) at room temperature.

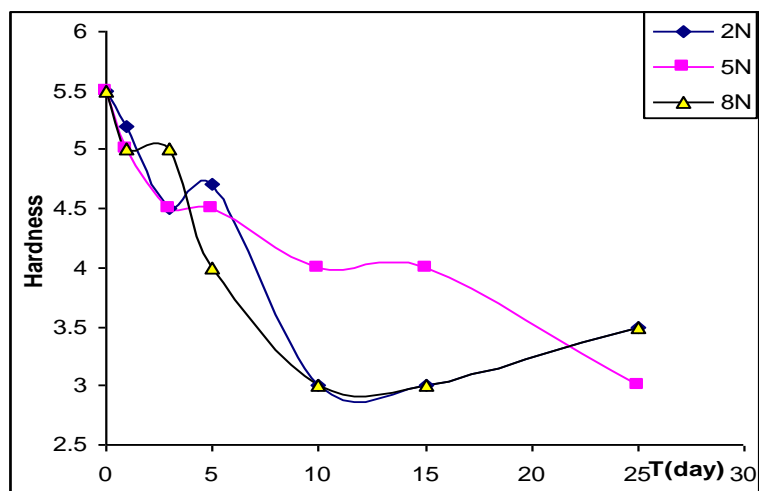


Fig. 13. Effect of immersion time (days) and HCl concentrations on the absolute hardness of glaze (G2) at room temperature.

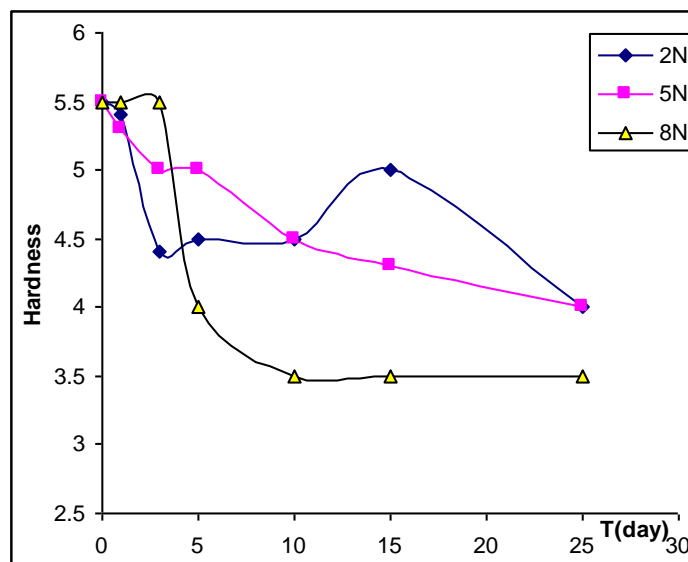


Fig. 14. Effect of immersion time (days) and HCl concentrations on the absolute hardness of glaze (G3) at room temperature.

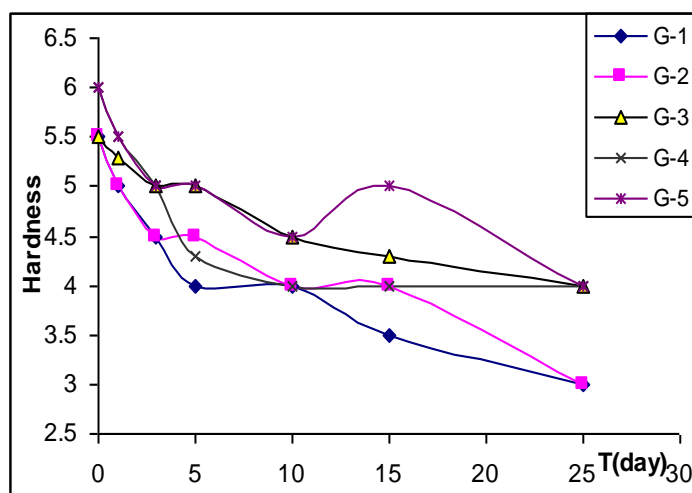


Fig. 15. Effect of glaze compositions and 5N HCl concentration on the absolute hardness at room temperature.

The leaching results (Fig.1-5) can be explained according to the postulations given in previous studies<sup>(10)</sup> which have shown that when sodium borosilicate glasses dissolve in acidic solutions, sodium and boron are selectively removed to produce hydrosilicate leached layers. In all cases, the structure of the leached glass is substantially different from that of the parent glass. It can be assumed that<sup>(16)</sup> some of the nonbridging oxygen (NBO) sites are converted into silanol (SiOH) groups via the  $H^+ \leftrightarrow Na^+$  ion exchange:



The large decrease in concentration of NBO's after ion exchange demonstrates that silanol groups in leached glass react with each other to form new Si-O-Si bonds via reaction such as:

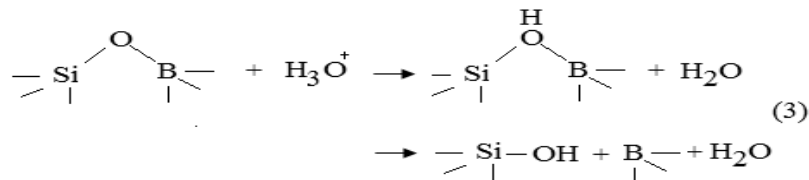


If reaction (2) was to go to completion, the product would be amorphous silica and water.

The leaching results clearly demonstrate that (CDGC) undergo some structural changes during leaching in high concentrated HCl aqueous solutions. Published results concerning the chemistry of silica indicate the formation of silanol groups during selective leaching promote the observed structural changes<sup>(17)</sup>. Silanol groups promote the hydrolysis of adjacent Si-O-Si bonds, leading to depolymerization and dissolution of the silicate network. Simultaneously silanol groups exhibit a strong tendency to condense with each other, reforming Si-O-Si bonds and promoting repolymerisation of the network causing false stability in durability observed in the interval times (5-15 days) indicating the repolymerisation of silicate network after leaching.

The fact that all the four glazes (G2-G5) have almost identical concentrations of both silanol groups and fourfold rings suggests that the structure of the hydrosilicate phases produced by leaching under acid conditions is independent of the structure of the parent glass. It appears that the initial network structure is broken down and reconstructed and at last time is completely dissociated (Fig.5) during leaching with high HCl concentration (8N). For borosilicate glasses, another important reaction leading to the formation of silanols is the hydrolysis of Si-O-B. It was suggested<sup>(16)</sup> that in acidic solutions hydrolysis of Si-O-B and B-O-B bonds (with either trigonal or tetrahedral borons) involves the electrophilic attack of protons on bridging oxygens (Reaction 3).

Reaction 3 explains why boron leaching is promoted in acidic solutions and why acidic solutions promote the conversion of fourfold boron sites to threefold sites. Dissolution data suggest that the pH below which the bridging oxygen is protonated and subsequently hydrolyzed for both Si-O-B and B-O-B bonds is near 4<sup>(16)</sup>.



The network is transformed from a homogeneous diffusion barrier to a porous aggregate colloidal silica particles in which diffusion is rapid and not contribute to the generation of large tensile stress in leached layers, leading to crazing and spalling<sup>(18)</sup>. However, increasing immersion time up to 25 days leads to fast and high dissolution due to the detach of leached layer (Fig. 4&5) and complete dissociation takes place for glazes G4 and G5.

The high dissolution observed with high temperature as represented in Fig. 6,7 and 8 can be explained as follows, at such high temperature (100°C), the convection currents promote ease ionic contacts and diffusion and ionization of further protons from the weakly dissociated acid replacing the consumed ones which leads to the observed continuous increase in corrosion and the high temperature promotes high dissolution.

The observed constancy or decrease in the weight loss when CKD exceeds 30% as observed in Fig. 8 can be realized and interpreted by assuming that the increase of CKD introduces much CaO and other constituents which may reach constant solubility with HCl and saturation with less ability for further attack.

#### *pH data*

The effect of HCl concentrations on the pH values either at room temperature or at 100°C is seen in Fig. 9 and 10, respectively. The pH appears to behave in a contradictory manner to that described for the dissolution. In other words, the high acidity of the solutions the greater the pH values (*i.e.* becomes more basic). This increase in pH values is assumed to increase the attack by HCl leading to high running and dissolution of alkalies from glazes to the leachant which form hydroxides causing the increase in pH values. Also, from these figures it is observed that the pH values increase when the CKD% exceeds 30% and this is due to the presence of K<sub>2</sub>O and CaO which form KOH and Ca(OH)<sub>2</sub>.

#### *SEM*

Micrographs of glaze surface (G3) before and after soaking in 2, 5 and 8N HCl solutions for 25 days have been shown in Fig. 11. The images are taken of the same surfaces and at identical locations of the surface line profile given in Fig.11. From these images grey holes are observed in the surface structure. After soaking in 2N HCl solution, these holes were partly dissolved and residual crystals were assumed to be nucleated. The crystals are seen in the 5N HCl surface as tiny peaks and after soaking, holes can be observed instead. After soaking in 8N HCl, clear changes in micro-roughness of the surface can be verified. This indicates that the increase in

micro-roughness caused by selective dissolution of the surface is the main reason for the observed differences in softness degree.

### *Hardness*

It had been reported<sup>(19)</sup> that, the glaze portion has a leached layer after basic (pH=12) dissolution but not after acidic (pH = 1) dissolution. Ion exchange during dissolution may cause the formation of M-OH and hence cracking of a leached layer<sup>(18,20)</sup>. Alternatively, the residual stress of the glaze due to devitrification and/or cooling may account for the cracking of a leached layer during dissolution. Drying may cause further cracking of this layer analogous to the case of gel drying<sup>(21,22)</sup>.

It is well known that silicate glasses are generally brittle<sup>(23)</sup>. Their stress intensity factor (SIF) is very low, but the actual situation is even worse for glass in ambient conditions, the tip of a surface crack undergoing a SIF of half the critical value, actually is slowly running. The crack speed does not depend on the environment; it increases sharply with SIF but levels off at some characteristic speed. As the strength of glass is controlled by the size of flaws, the fatigue effect is related to the growth of cracks (stress corrosion) during aging in the given environment, under load. Also, the stress corrosion issue requires a careful control of the residual stress.

Our results (Fig. 12 -14) demonstrate that there is a decrease in the hardness values in the early periods of immersion times (up to 10 days) followed by a stability in the hardness values. This behavior is due to the progressive alkaline leaching out of the glass, and that this change in the chemical composition at the tip of the crack is responsible for the fatigue limit rather than a geometrical change (blunting).

The constancy of hardness values (Fig. 12 - 14) after long immersion time (10-25 days), can be assumed to the formation of silanol groups which can be associated with depolymerization and repolymerisation reactions involving the silicate network.

It was suggested that<sup>(16)</sup>, the extent of network repolymerization and the nature of the leached glass structure are controlled in large part by the concentration and distribution of silanol groups created in the glass during leaching. Then, it is possible to make predictions concerning how the glass composition, structure and solution chemistry should influence the repolymerisation process and leaching kinetics. In addition, the ease with which the network will repolymerize should increase as the number of silanols created during leaching increases. No polymerization should be observed for fused silica and glasses with a high silica content should be resistant to repolymerisation while alkali silicate with a high alkali content should be very susceptible to repolymerisation. Moreover, alkali aluminosilicate or borosilicate should not undergo repolymerisation in the neutral pH regime, but should undergo repolymerisation below pH 4 where selective leaching of Al or B leads to the formation of silanols and this is our case.



Referring to Fig. 15, showing the successive increase in hardness (G1→G5) and this increase is relative to the amount of CKD added which is responsible for the hardness characteristic of the glaze due to its high content of CaO, MgO and Fe<sub>2</sub>O<sub>3</sub>. In addition, the composition of the glazes in these samples has relatively high content of zircon and dolomite.

### Conclusion

It can be concluded that CKD waste has potential to be used in the preparation of new commercial industrial glazes. The amount of CKD utilized in the glaze formulation can reach up to 50%. Our results indicate that, the CKD affect the melting temperature, increasing CKD% lowers the melting temperature. The chemical corrosion is markedly degraded due to high concentrated HCl (2,5&8N), but even with these concentrations the new prepared glazes exhibit resistance for these concentrations especially G2 and G3 and improvement of hardness is relative to CKD%. Finally, experiments on the CDGC enable the extent to which the CKD matrix affects mineral dissolution to be assessed. A leached layer is assumed to be produced under extreme acidic dissolution (8N = 8N) of the present glaze, because a high dissolution may interfere between both the adsorption of cations and the reconstruction of the surface.

Hence based on these observations, the CKD glazes are suitable to be used for glazing wall tiles which are not subjected to high pressure. In addition, these glazes can also be used for glazing table wares. The possibility of using glazes with a high CKD content (10-50%) facilitates the use of cheap materials containing calcium and iron compounds and carbonates.

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## دراسة التآكل للجليز السيراميكي المخلوط بتراب الاسمنت المغمور في حامض الهيدروكلوريك المركز

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تهدف الدراسة الى قياس مقدار التآكل الناتج من عينات الجليز السيراميكي المخلوط بنسب مختلفة من تراب الاسمنت (10 : 50%) ومقارنتها بالجليز السيراميكي التجارى ، وقد وجد ان مقدار التآكل فى هذه العينات يزيد بزيادة تركيز تراب الاسمنت كذلك تمت دراسة تأثير التركيزات العالية المختلفة من حامض الهيدروكلوريك المركز (8,5,2 عيارى) كذلك تمت دراسة التآكل عند درجة حرارة عالية 100 م ووجد ان التآكل ايضا يزداد عند الحرارة العالية. وقد تم شرح النتائج على نظرية التبادل الايونى مما يؤدى الى ذوبان ملحوظ فى سطح الجليز يتوقف على قوة المحلول وزمن الغمر. كذلك تم قياس الأس الهيدروجيني لمحاليل التآكل ووجد أنه يقل مع زيادة التآكل وذلك لتكون أحماض السيليكات والبوريك بالإضافة الى ارتفاع تركيز محلول الغمر. وقد تم ايضا تصوير سطح العينات الزجاجية بعد تأكلها بواسطة الميكروسكوب الالكترونى لتوضيح شكل السطح بعد التآكل بالمحاليل.

ايضا تم دراسة الصلادة للعينات المحضرة قبل وبعد الغمر فى المحاليل وقد وجد أن الصلادة تقل عند غمر العينات فى محاليل التآكل ولكن تظل محتقظة بقيمة الصلادة المسموح بها فى الصناعة ، كذلك بمقارنة الصلادة للعينات وجد ان الصلادة تتحسن بزيادة نسبة تراب الاسمنت وذلك لإرتفاع نسبة الكالسيوم فى تراب الاسمنت.

إن من أهم أهداف هذه الدراسة هو المساهمة فى حل المشكلات البيئية الناجمة عن المخلفات الصناعية مثل تراب الاسمنت كذلك فإنها تفتح المجال لكثير من الدراسات التى تهدف الى تدوير المخلفات الصناعية التى تحتوى على نسبة كبيرة من الفلويات والحديد فى انتاج انواع جديدة من الجليز السيراميكي الذى يستخدم فى الصناعة وقد وجد انه عند اضافة التراب بنسبة تصل الى 30% فان الجليز الناتج يعطى خصائص جيدة مما يسمح بامكانية تطبيق ذلك عمليا.