

Comparative Study on Synthesis and Characterization of Ceria Based Composite Oxides Containing Manganese Nano-Particles

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M_n/Ce nano-composite oxides containing different amounts of manganese (5, 9, 17, 23, 29, 34 wt% as MnO₂) were prepared by impregnation method at various calcination temperatures. Effect of both thermal treatment and loading on the structural, surface and microstructure properties of the as prepared nano-composites was determined. The combined effect of manganese oxide and ceria, at different concentrations, strongly influences the previous properties of the nano-composite oxides, by dispersing the MnO₂ phase and promoting the efficiency of the Ce⁴⁺- Ce³⁺ and Mn⁴⁺ - Mn³⁺ redox couples.

The thermal treatment and loading influenced the interaction between manganese and CeO₂ evidently. The incorporation of Mn ions into CeO₂ crystal lattice resulted in weaker interaction between manganese and ceria on composite surface. Manganese loading at 400°C led to a slight increase in the ceria crystallite size which decreased by increasing the calcination temperature from 400 to 600°C. The sintering activation energy of ceria was evaluated to be 12kJ/mol for the MnO₂/CeO₂ nano-composite.

Keywords: XRD, TEM, S_{BET} and MnO₂/CeO₂ nano-composite.

Most oxide ceramics, especially cerium oxide, are known to play an important role in many catalytic industrial processes ⁽¹⁾. The physicochemical properties of these oxides and their solid solution derivatives are greatly dependent upon three main factors: particle size, structural distortion and chemical nonstoichiometry ⁽²⁾. In general, reducing the particle size of a catalyst solid results in increasing its surface area and changing its morphology (*e.g.*, from large pellets to spherical clusters), thus providing a larger number of more reactive edge sites. Especially when the particle size is decreased below 100 nm, the materials become nano-phase where the density of defects increases so that up to 50% of the atoms (molecules) are situated in the cores of defects (grain boundaries, inter-phase boundaries, dislocations, ... etc.). The high density of the defects in nano-phase materials provides a large number of active sites for gas-solid catalysis, while the

high diffusivity through nanometer-sized inter-phase boundaries promotes fast kinetics of catalyst activation and reactions. Stoichiometric defects such as oxygen deficiency (vacancy) and its mobility on the catalyst surface are of great importance for redox gas conversions. Oxygen vacancies can be created by non-stoichiometric synthesis of the oxides of a metal which has either multiple oxidation states or partial substitution (doping) of another cation with lower oxidation state. The mobility of the oxygen vacancies increases with decreasing particle size and increasing the inter-granular layers.

Cerium dioxide, CeO_2 , is known to play an important role in many catalytic processes⁽²⁾. CeO_2 has a fluorite-like cubic structure in which each cerium site is surrounded by eight oxygen sites in fcc arrangement and each oxygen site has a tetrahedron of cerium sites. Under various redox conditions, the oxidation state of cerium may vary between 3 + and 4 + .

Ceria-based mixed oxides ($\text{M}_x\text{Ce}_{1-x}\text{O}$) are versatile solid oxygen exchangers. At temperatures in the range of 400–800 °C, the redox couple $\text{Ce}^{4+} \leftrightarrow \text{Ce}^{3+}$ facilitates oxygen storage and release from its bulk fluorite lattice, making it an ideal candidate for catalytic and/or electrocatalytic oxidation applications in solid oxide fuel cells (SOFCs). However, the surface redox chemistry of ceria is sensitive even at low temperatures to crystal structure defects, which can be tuned by substituting some of the Ce cations with ions of different sizes and/or charges. Substitution of a lower-valent metal ion (*e.g.*, M^{3+}) by cerium lowers the energy barrier for oxygen migration⁽³⁾. However, smaller homovalent ions (M^{4+}) enhance the oxygen storage capacity (OSC) by decreasing the activation energy for the reduction ($\text{Ce}^{4+} \rightarrow \text{Ce}^{3+}$) and retarding OSC degradation at high temperature. Given the effects of trivalent ions and smaller sizes on the structure and properties, there is considerable scientific interest in introducing M^{3+} ions into the ceria lattice. Ceramic methods allow the formation of such solid solutions, but the preparation temperature and time required to obtain a homogeneous mixture depends to a significant extent on the particle size of the starting material. The different properties of some oxide ceramics and their solid solution derivatives are greatly dependent upon three main factors: particle size, structural distortion and chemical non-stoichiometry⁽²⁾.

It is noted that the composition of Mn-Ce catalysts exerts a strong effect on their redox properties at low temperature^(3,4). Transition metal Mn oxides with variable valence states can store and release oxygen in the same way as CeO_2 and Ce_2O_3 ; therefore, they may be oxygen storage materials, which are main components for automotive exhaust catalyst. It is possible for Ce-Mn-O composite oxides to be a great potential as additives to three- way catalyst.

The mixed Mn-Ce-O oxides were studied in many interesting fields. Mn-Ce composite oxides have been found to be remarkably active in the sub-critical wet oxidation of several refractory pollutants, such as ammonia, acetic acid, pyridine, phenol and polyethylene glycol⁽⁵⁾. It is well established that the preparation

method could exert a basic influence on the physicochemical properties of ceria-based mixed oxides, like surface area, component dispersion and strength of interaction, which in turn determines the redox properties and reactivity of the final product ⁽⁶⁾. It has been well accepted that the interactions between MnO_x and CeO₂ vary with composition, resulting in the evolution of textural, structural, and oxidation state ⁽⁷⁾.

Manganese–cerium composite oxides with different molar ratio Mn/ (Mn + Ce) (0, 0.25, 0.50, 0.75, 1) were prepared by citric acid method and investigated concerning their adsorption behaviour, redox properties and behaviour in the selective catalytic reduction of NO_x by NH₃ ⁽⁴⁾. The studies based on pulse thermal analysis combined with mass spectroscopy and FT-IR spectroscopy uncovered a clear correlation between the dependence of these properties and the mixed oxide composition. The highest activity to nitrogen formation was found for composites with a molar ratio Mn/(Mn + Ce) of 0.25, whereas the activity was much lower for the pure constituent oxides. Yasyerli *et al.* ⁽⁸⁾ argued for the formation of Mn-Ce-O solid solutions and observed shift in the lattice parameter of ceria on cerium–manganese composite oxides containing different molar ratios of Ce/Mn. They concluded that the high reducibility and catalytic activity of the mixed oxides were associated with strong interactions between Mn and Ce.

In order to improve the structural and morphological properties of these materials, we used the impregnation method to prepare new samples with varying Mn/Ce ratio in the entire compositional range at different calcination temperatures. This procedure allows an excellent control over relevant properties of the synthesized materials, namely compositional homogeneity, purity, and surface area. In this work, the samples were characterized by different techniques.

Experimental

Composite preparation

Various MnO₂/CeO₂ composite samples were prepared by impregnating a known weight of finely powdered ceria solid with calculated amounts of manganese dissolved in the least amount of distilled water enough to make a paste. The paste containing manganese was dried at 100 °C until constant weight and was then subjected to heat treatment for 4 hr at 400 °C, respectively. The manganese concentrations were 5, 9, 17, 23, 29 and 34 wt% as MnO₂. Additional samples were prepared by heating the paste containing 17wt% MnO₂ for 4 hr at 500 and 600 °C. The chemicals employed were of analytical grade and supplied by BDH Company.

Composite characterization

An X-ray measurement of various mixed solids was carried out using a BRUKER D8 advance diffractometer. The patterns were run with Cu K_α radiation at 40 kV and 40 mA with scanning speed in 2θ of 2 ° min⁻¹.

The crystallite size of CeO₂ phase present in the investigated solids was based on X-ray diffraction line broadening and calculated by using Scherrer equation⁽⁹⁾.

$$D = \frac{B\lambda}{\beta \cos \theta} \quad (1)$$

where D is the mean crystallite size of the phases under investigation, B is the Scherrer constant (0.89), λ is the wave length of X-ray beam used, β is the full-width half maximum (FWHM) of diffraction and θ is the Bragg's angle.

Transmission electron micrographs (TEM) were recorded on JEOL TEM-1230 electron microanalysis. The samples were dispersed in ethanol and then treated ultrasonically in order to disperse individual particles over a gold grids.

Surface characteristics

The surface characteristics of various solid catalysts, namely, the specific surface area (S_{BET}), total pore volume (V_p) and mean pore radius (\bar{r}) were determined from nitrogen adsorption isotherms measured at -196 °C, using Nova 2000, Quanta Chrome (commercial BET unit). Before undertaking such measurements, each sample was degassed under a reduced pressure of 10^{-5} Torr for 2 hr at 200 °C.

Results

Structural analysis

XRD examination of MnO₂/CeO₂ composites, containing 5, 9, 17, 23, 29 and 34 wt% MnO₂, calcined at 400 °C for 4 hr revealed a nano-crystalline nature for the bulk structure. The crystalline phase compositions identified are shown in Fig. 1. Inspection of this figure showed the following: (i) the diffraction patterns of each sample evidenced the nanocrystalline CeO₂ phase was observed. The ceria is cubic system with space group *Fm3m*. The diffraction of the 4-29 wt % MnO₂/CeO₂ samples calcined at 400 °C includes all the diffraction peaks of CeO₂ with the absence of MnO₂ peaks depending on the amount of MnO₂ in these samples which is very small beyond the detection limit of X-ray diffraction. This observation suggests fine dispersion of MnO₂ on the surface of ceria and/or Mn-Ce-O solid solution formation⁽⁸⁾. (ii) The X-ray diffractograms of 34 wt % MnO₂/CeO₂ sample subjected to heating at 400 °C showed that this sample consisted of all diffraction peaks of CeO₂ as a major phase besides some diffraction peaks of poorly crystalline MnO₂ (Pyrolusite) as a minor phase⁽¹⁰⁾. (iii) MnO₂ loading of the investigated system calcined at 400 °C brought about some modifications in the degree of crystallinity of the produced phases. The intensity of the (111) plane of CeO₂ was considered as a measure of its degree of crystallinity. Manganese loading resulted in a progressive decrease in the degree of crystallinity of CeO₂ phase as evidenced by the significant decrease in the height of all diffraction peaks of this oxide and also an increase in their width.

This decrease was more pronounced in case of the 23 wt% $\text{MnO}_2/\text{CeO}_2$ sample. The maximum decrease in the intensity of the (111) plane of CeO_2 due to loading by 23 wt% MnO_2 was attained 74%.

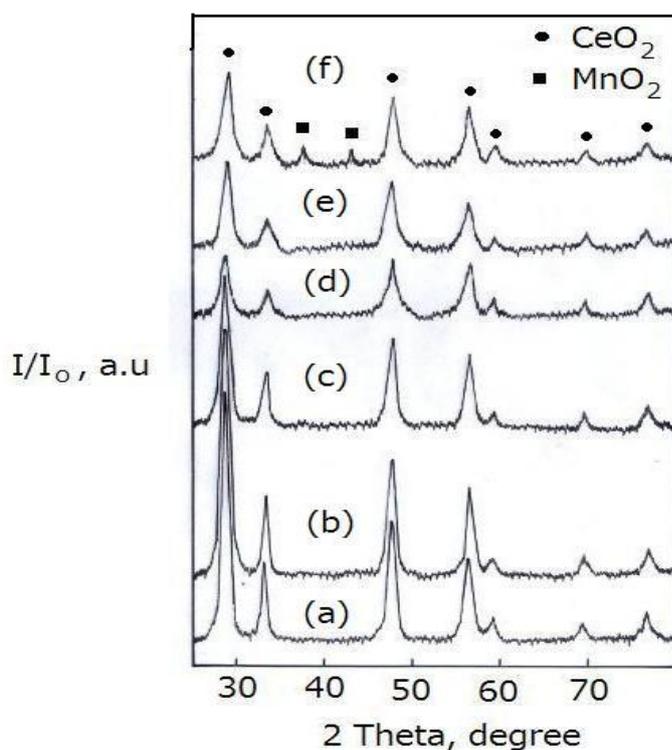


Fig. 1. X-ray diffractograms of $\text{MnO}_2/\text{CeO}_2$ nanocomposites calcined at $400\text{ }^\circ\text{C}$. MnO_2 concentrations: (a) 5; (b) 9; (c) 17; (d) 23; (e) 29 and (f) 34 wt%.

The effects of the calcination temperature on the X-ray patterns of the 17 wt % $\text{MnO}_2/\text{CeO}_2$ sample were investigated. Figure 2 displays the X-ray diffractogram of the 17 wt % $\text{MnO}_2/\text{CeO}_2$ sample calcined at 500 and $600\text{ }^\circ\text{C}$ for 4 hr. The comparison between Fig. 1 and 2 shows the following: (i) the XRD patterns of the 17 wt % $\text{MnO}_2/\text{CeO}_2$ sample consisted only of all diffraction peaks of CeO_2 . (ii) The rise in the calcination temperature from 400 to $600\text{ }^\circ\text{C}$ led to a decrease in the degree of crystallinity of CeO_2 phase as evidenced by the significant decrease in the height of all diffraction peaks of this oxide and also an increase in their width. The maximum decrease in the intensity of the (111) plane of CeO_2 due to heating of the 17 wt % $\text{MnO}_2/\text{CeO}_2$ sample calcined at $600\text{ }^\circ\text{C}$ attained 62.5%. This indicates that the heat treatment enhanced the solubility of manganese oxide in the ceria lattices.

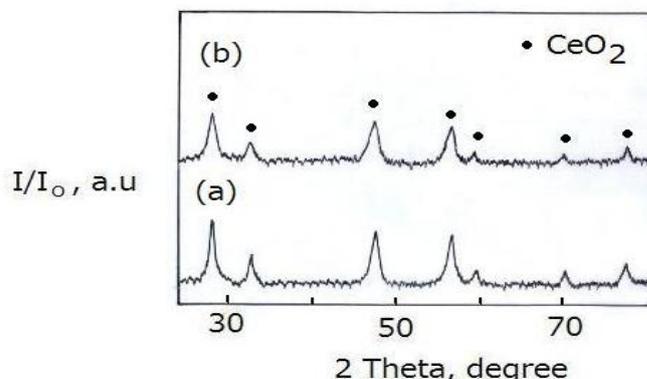


Fig. 2. X-ray diffractograms of 17wt% MnO₂/CeO₂ nanocomposites calcined at (a) 500 and (b) 600 °C.

The X-ray data enabled us to investigate the effects of the calcination temperature and MnO₂ loading of the system studied in modifying the lattice constant, crystal size, unit cell volume and degree of crystallinity of CeO₂ phase present in various composites. The computed values of different structural parameters of CeO₂ phase in different composites are given in Table 1. The careful examination of this table reveals the following: (i) MnO₂ loading of the system investigated carried out at 400 °C resulted in a measurable increase in the crystallite size, lattice constant and unit cell volume of CeO₂ phase depending upon the amount of MnO₂ added. The addition of 23 wt% MnO₂ to the system investigated brought about an increase of 19.2, 1.7 and 0.6 % in the crystallite size, lattice constant and unit cell volume of CeO₂ phase, respectively. (ii) The augmentation in calcination temperature of the 17 wt % MnO₂/CeO₂ composite from 400 to 600 °C brought about a significant decrease in the crystallite size, lattice constant and unit cell volume of CeO₂ phase. The maximum decrease attained 38.5, 1.5 and 0.4 % in the crystallite size, lattice constant and unit cell volume of CeO₂ phase, respectively, due to heating the 17 wt % MnO₂/CeO₂ composite at 600 °C. (iii) The values of crystallite size of CeO₂ crystallite present in the investigated composites calcined at 400–600 °C enabled the calculation of sintering activation energy (ΔE_s) of this oxide using Arrhenius equation, this was found to be 12 kJ/mol⁽¹¹⁾.

The observed structural modifications of CeO₂ phase due to heat treatment and/or loading of MnO₂/CeO₂ system with different amounts of MnO₂ can be discussed in terms of dissolution of some manganese species in CeO₂ lattice. The dissolution process may lead to the formation of new chemical compounds and/or solid solutions⁽¹⁰⁾. The dissolution of MnO₂ in the outermost surface of ceria led to conversion of some Ce⁴⁺ ions, present in non-stoichiometric CeO₂, into Ce³⁺ ions with subsequent increase in the value of its lattice constant. On the contrary, the inhibiting effect might be due to the diffusion of MnO₂ into the interior of ceria grains by the heat treatment.

TABLE 1. The effects of heat treatment and MnO₂ loading on some structural parameters of cerium oxide involved in MnO₂/CeO₂ nanocomposite.

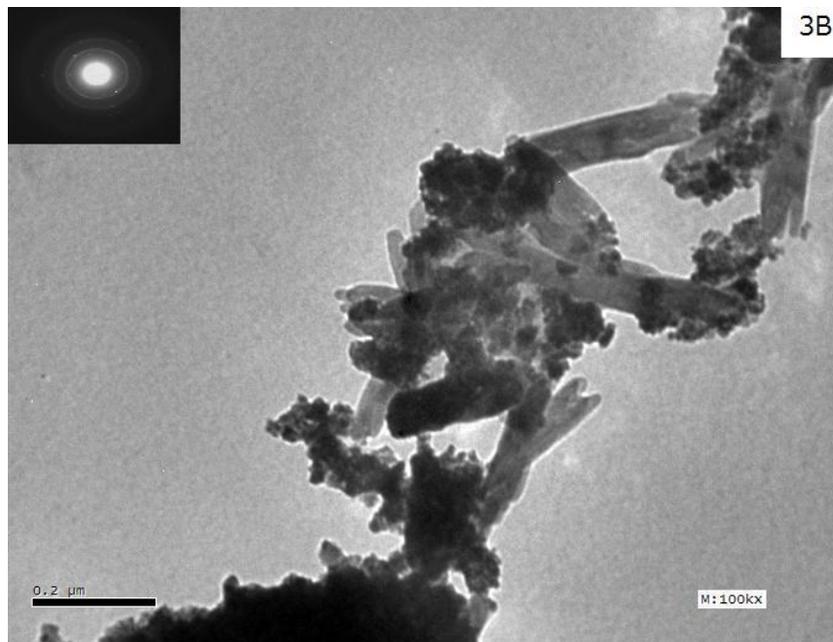
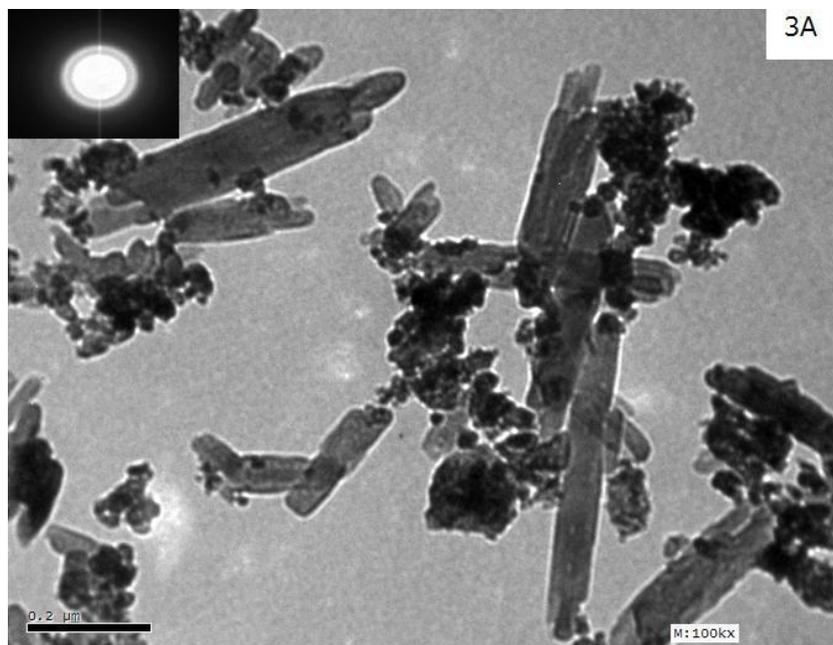
Concentration of MnO ₂ (wt%)	Calcin. Temp. (°C)	D (nm)	a (nm)	V (nm ³)	Degree of crystallinity (a.u)	Crystalline composition ^a
5	400	15.6	0.5403	0.1577	59	CeO ₂
9	400	16	0.5408	0.1582	57	CeO ₂
17	400	17.9	0.5414	0.1587	40	CeO ₂
	500	15	0.5409	0.1583	17	CeO ₂
	600	11	0.5406	0.1580	15	CeO ₂
23	400	18.6	0.5434	0.1605	16	CeO ₂
29	400	16.8	0.5414	0.1587	28	CeO ₂
34	400	16	0.5412	0.1586	33	CeO ₂ (d, j) and MnO ₂ (m)

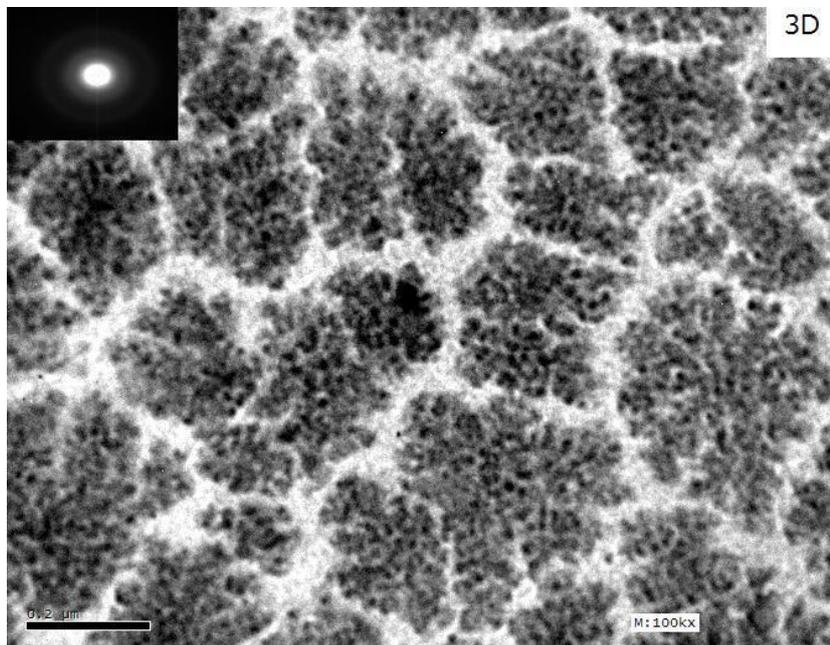
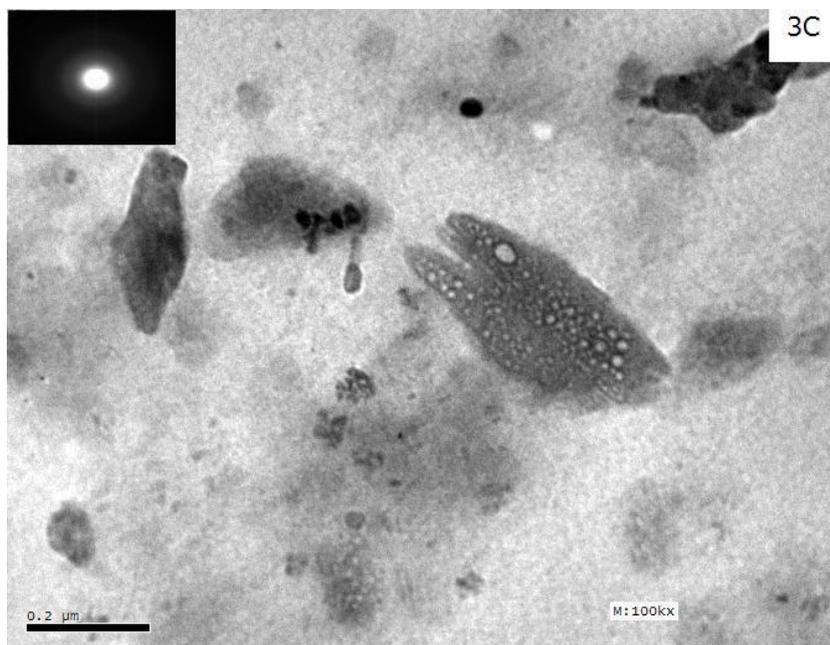
^a d: dominant, j: major, m: minor.

TEM analysis

The micro structural evolution during the heat treatment and loading of MnO₂/CeO₂ composite samples are shown in Fig. 3. Figure (3A-C) shows the effect of loading by 5, 17 and 34 wt % MnO₂ on the formation of MnO₂/CeO₂ composite particles under impregnation conditions. It shows agglomeration of ultra fine particles with rod-like structures. These samples are revealed to be assembled by randomly oriented spherical nano-crystallites. The population size distribution calculated from TEM images shows the particle diameter for MnO₂/CeO₂ composites calcined at 400 °C of about 15-18 nm. When the ceria has been modified with increasing amounts of manganese oxide, we can find that the surface of spherical crystallites is covered very uniformly with highly ordered array. However, the manganese loading resulted in a slight increase in the particle size of the investigated solids indicating the formation of Mn-Ce-O solid solution at the composite surface. Fig. 3(D & E) displays the effect of heat treatment at 500 and 600°C on the morphology of loading by 17 wt % MnO₂/CeO₂ composite systems.

The comparison between TEM images in Fig. 3(B, D) and E shows the following: (i) the rise in the calcination temperature from 400 to 500 °C brought about tree leaves like particle composed of spherical and uniform crystallites. Increasing the preparation temperature from 500 to 600°C led to weak agglomeration of the nanoparticles. (ii) The heat treatment resulted in a decrease in the particle size of the investigated composite. In fact, the solid calcined 600°C as shown in Fig. 3E showed the smallest crystallite size of diameter about 11 nm. In other words, the heat treatment and MnO₂ loading brought about a decrease in the particle size of MnO₂/CeO₂ composite. These findings suggest incorporation of manganese species in the ceria lattices indicating the contraction in the ceria lattice with subsequent decrease in the solid state reaction between MnO₂ and CeO₂ forming Mn-Ce-O solid solution at the composite surface. Moreover, the stimulation effect towards the solubility of manganese species in the ceria crystal was more pronounced by heat treatment.





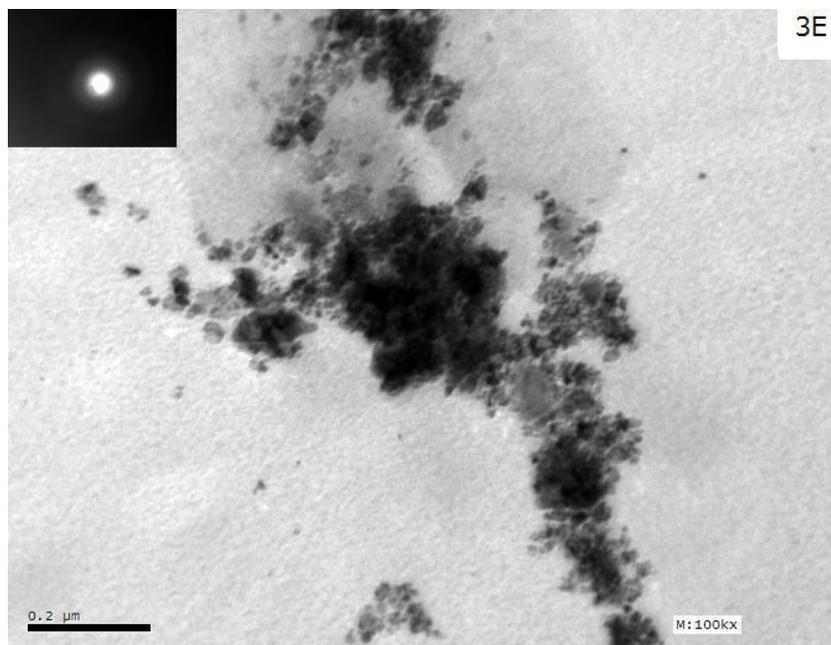


Fig. 3. TEM images of various $\text{MnO}_2/\text{CeO}_2$ nanocomposites: (A) 5 wt% $\text{MnO}_2/\text{CeO}_2$ calcined at 400 °C, (B) 17 wt% $\text{MnO}_2/\text{CeO}_2$ calcined at 400 °C, (C) 34 wt% $\text{MnO}_2/\text{CeO}_2$ calcined at 400 °C, (D) 17 wt% $\text{MnO}_2/\text{CeO}_2$ calcined at 500 °C and (E) 17 wt% $\text{MnO}_2/\text{CeO}_2$ calcined at 600 °C.

To complete the information, the electronic micro-diffraction was used to determine the chemical structure of the small grains. The corresponding selected area electron diffraction (SAED) patterns of the mentioned samples are shown in Fig. 3(A-E). All the diffraction spots corresponding to aggregates are generally distributed on circles. Whatever the sample, the circles related to the micro-diffraction of CeO_2 are continuous in agreement with the homogeneity of the CeO_2 grains observed in TEM images. The ring-type diffraction pattern shown in Fig. 3(A-E) is indexed to polycrystalline CeO_2 in fluorite structure.

Surface Characteristics

The surface properties of the various investigated solids determined from nitrogen adsorption isotherm measured at -196 °C are given in Table 2. The data in this table demonstrate the effects of manganese content and calcination temperature on the different surface properties of the investigated mixed solids. Representative adsorption/desorption isotherms for nitrogen on 17 wt% $\text{MnO}_2/\text{CeO}_2$ composite sample calcined at 400-600 °C are depicted in Fig. 4. All the isotherms displayed common characteristics, being similar in the shape to type II of Brunauer's classification⁽¹²⁾.

TABLE 2 . The surface properties of Mn/Ce nano-composite catalysts containing different amounts of manganese oxide.

Concentration of MnO ₂ (wt%)	Calcin. Temp. (°C)	S _{BET} (m ² /g)	V _p (cm ³ /g)	(nm) ^r
5	400	34	0.0768	1.91
9	400	40	0.0869	1.91
17	400	46	0.0877	1.91
	500	35	0.0603	1.91
	600	29	0.0557	1.91
23	400	52	0.0881	1.91
29	400	57	0.0899	1.91
34	400	61	0.0909	1.91

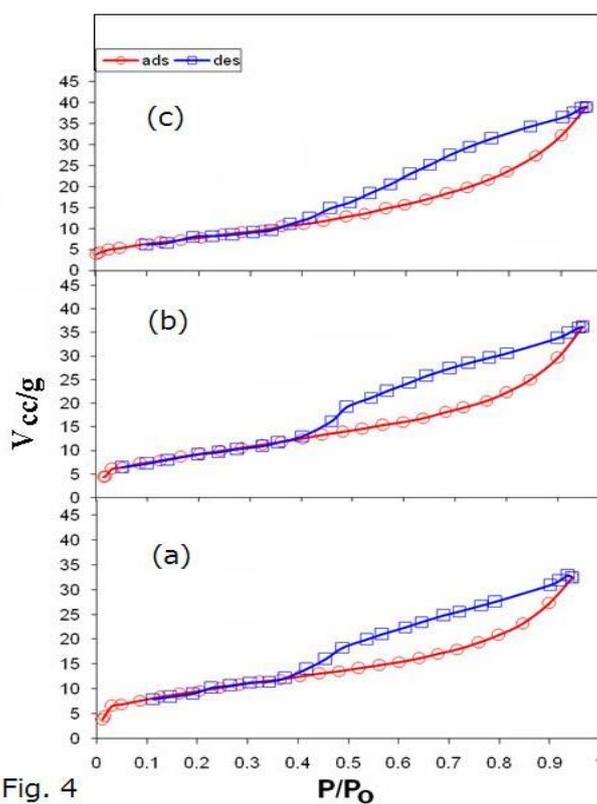


Fig. 4. N₂ adsorption/desorption isotherms for 17wt% MnO₂/CeO₂ nanocomposites calcined at (a) 400, (b) 500 and (c) 600 °C.

It is seen from Table 2 that the S_{BET} values for 17 wt% $\text{MnO}_2/\text{CeO}_2$ composite decreased as the calcination temperature was increased from 400 to 600 °C. However, the loading process resulted in an increase (79.4%) in the S_{BET} value of Mn/Ce nanocomposite oxides calcined at 400 °C.

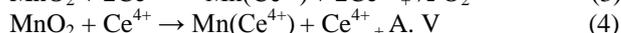
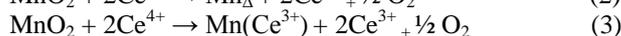
The increase in the S_{BET} values for the investigated oxides calcined at 400 °C may be explained in terms of the following: (i) The creation of new pores resulting from the liberation of gaseous nitrogen oxides during the thermal dissociation of manganese and cerium nitrates⁽¹³⁾. (ii) The augmentation in the total pore volume of Mn/Ce nanocomposite oxides.

The decrease in the S_{BET} values for 17 wt% $\text{MnO}_2/\text{CeO}_2$ composite calcined in air at 400- 600 °C may be discussed in terms of the following: (i) Sintering of the mixed oxides due to the aggregation of Mn/Ce ions via increasing of their mobility⁽¹³⁾. (ii) The decrease in the total pore volume of Mn/Ce nanocomposite oxides.

Discussion

The addition of manganese oxides to the ceria surface resulted in some modifications in the structural, surface and morphological properties of the final product. At low metal content, metal species diffuse into the fluorite forming a solid solution with ceria and also metal oxide highly dispersed on CeO_2 surface. At high metal content, solid solution, crystallized metal oxide and highly dispersed metal oxide phases on ceria are co-existing. The ability of ceria based materials to reversibly accept or contribute oxygen to its surroundings is the basis for its use in catalytic applications, where redox processes are important. The exact nature of the activity of cerium-based oxides depends primarily on the nature and concentration of defects in the material (*e.g.* oxygen vacancies, Ce^{3+} centres and dopants). When oxygen vacancies are present, oxygen ions can move through the lattice relatively easily giving rise to high-oxygen ion conductivity⁽¹⁴⁾. This also plays an important role in catalytic applications because oxygen vacancies can move rapidly from surface to bulk and vice versa. Identifying the most active phases and understanding the phase changes taking place under a certain treatments is critical to develop improved catalysts.

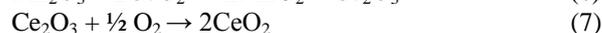
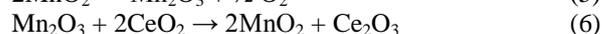
The observed structural modifications of CeO_2 phase due to loading of ceria based system with different amounts of manganese oxides can be discussed in terms of dissolution of some of these oxides in CeO_2 lattice. The dissolution process may lead to the formation of new chemical compounds and/or solid solutions⁽¹¹⁾. Ceria can dissolve various amounts of MnO_2 added. The dissolution of these oxides can take place via substituting some of host cerium ions in cerium oxide lattice by the cations added. The dissolution process in case of MnO_2 loading can be simplified adopting Kroger's notions according to⁽¹⁵⁾:



Egypt. J. Chem. **53**, No.1 (2010)

where Ce^{3+} denotes to trivalent cerium ion present in nonstoichiometric CeO_2 lattice as lattice defects; $\text{Mn}(\text{Ce}^{3+})$ and $\text{Mn}(\text{Ce}^{4+})$ are tri- and tetravalent manganese ion located in the position of tetravalent and trivalent cerium ions of cerium oxide, receptivity; Mn_Δ is manganese ion located in the interstitial positions of cerium oxide lattices; A.V. created anionic vacancies. The incorporation of Mn^{4+} in the lattice of CeO_2 according to the mechanism represented by reactions (2) and (3) is accompanied by conversion of some tetravalent cerium ions into trivalent cerium ions with subsequent removal of some of excess oxygen present in non-stoichiometric CeO_2 solid leaving oxygen vacancies. This observed, experimentally, via increasing of Mn content resulted in an increase in the total pore volume of the mixed solids calcined at 400 °C indicating the increase in their surface area.

So, loading of cerium oxide by MnO_2 taking place according to reactions (2) and (3) should be accompanied by a decrease in the oxidation state of cerium ion in CeO_2 via conversion of some of Ce^{4+} into Ce^{3+} with subsequent expansion in the ceria lattice. This process is normally accompanied by an increase in the lattice constant of CeO_2 which has been found experimentally. In fact, MnO_2 loading at 400 °C resulted in a slight increase in the particle size of ceria solid. In addition, the MnO_x deposits coexist in intimate contact with CeO_2 crystallites, favouring oxygen transfer between the two metal oxides. This type of $\text{MnO}_x/\text{CeO}_2$ structure interaction led to an increase in the formation of the Ce_2O_3 phase. This indicated that there existed a synergistic mechanism between the manganese and cerium oxides. The proposed mechanism in this case is as follows:



It is well known that Ce_2O_3 can be easily re-oxidized to CeO_2 by exposure to air. On the other hand, the dissolution of manganese ions in the lattices of reacting oxides according to reaction (4) which led to creation of anionic vacancies might increase the mobility of cations of reacting oxides thus enhancing the formation of Mn-Ce-O solid solution. But, the incorporation of Mn species in the ceria lattice brought about a decrease in the ceria lattice parameter with subsequent decrease in the particle size due to the smaller ionic radius of Mn than that of Ce ions. Indeed, the heat treatment of 17 wt% $\text{MnO}_2/\text{CeO}_2$ composite sample at 600 °C led to a decrease in the particle size of ceria crystallites. However, the diffraction lines related the manganese phase were undetected in the XRD patterns for all samples studied except presence of trace amount of MnO_2 phase in case of the 34 wt% $\text{MnO}_2/\text{CeO}_2$ sample.

These observations suggest that MnO_2 loading at 400 °C led to a formation of solid solution between the manganese and cerium oxides at the composite surface, while an increase of the calcination temperature from 400 to 600 °C brought about incorporation of manganese species in the ceria lattice with subsequent migration of these species from the surface of composite to its bulks.

Conclusions

The following are the main conclusions that may be drawn from the results obtained in the present work:

1. Ceria based composites containing different amounts of manganese oxide prepared by impregnation method and calcined at 400 - 600 °C consisted of nanocrystalline solids as shown in XRD and TEM measurements.
2. Increasing the amount of manganese oxide brought about a decrease in the degree of crystallinity with subsequent increase in particle size, and lattice parameter, unit cell volume of ceria over Mn/Ce composite oxides indicating the formation of Mn-Ce-O solid solution via interaction between the manganese species and ceria.
3. The heat treatment of the MnO₂/CeO₂ composite resulted in a decrease in the degree of crystallinity, particle size, and lattice parameter, unit cell volume of ceria present.
4. The sintering activation energy of the MnO₂/CeO₂ composite was found to be 12 kJ/mol.
5. The heat treatment led to a decrease in S_{BET} of Mn/Ce composite oxides while an increase of the manganese content brought about an increase in the S_{BET} value.

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مقارنة على تحضير وتوصيف اكاسيد كومبوزيت معتمدة على السيريا والمحتوية على جزيئات المنجنيز النانوية

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تم تحضير اكاسيد نانوكومبوزيتية من المنجنيز والسيريوم محتوية على كميات مختلفة من المنجنيز (5 و 9 و 17 و 23 و 29 و 34٪ وزنا اكسيد منجنيز) باستخدام طريقة النقع عند درجات تحميص مختلفة. وتم دراسة تأثير المعالجة الحرارية و عملية التحميل بأكسيد المنجنيز على الخواص التركيبية والنسجية والشكلية للنانوكومبوزيتات المحضرة ووجد ان اتحاد أكاسيد المنجنيز والسيريوم- عند تركيزات مختلفة يؤثر بقوة على الخواص السالفة الذكر لأكاسيد النانوكومبوزيت من خلال انتشار اكسيد المنجنيز على سطح أكسيد السيريوم وتنشيط كفاءة ازواج كاتيونات المنجنيز والسيريوم ($Ce^{4+}-Ce^{3+}$ and $Mn^{4+}-Mn^{3+}$). لقد تأثرت كل من المعالجة الحرارية و عملية التحميل بالمنجنيز على التفاعل بين اكاسيد المنجنيز والسيريوم كما أن ادخال ايونات المنجنيز فى الشبكة البلورية لأكسيد السيريوم ادى الى ضعف التفاعل بين المنجنيز والسيريوم الموجود عند سطح الكومبوزيت . علاوة على أن عملية تحميل كميات متزايدة من المنجنيز عند درجة حرارة 400 درجة مئوية أدت الى زيادة طفيفة فى الحجم البلورى لأكسيد السيريوم الذى ينقص بدوره من خلال زيادة درجة حرارة التحميص من 400 الى 600 درجة مئوية وكانت قيمة طاقة تنشيط عملية التلييد 12 كيلو جول لأكسيد السيريوم الموجود فى نانوكومبوزيت أكاسيد المنجنيز والسيريوم.

